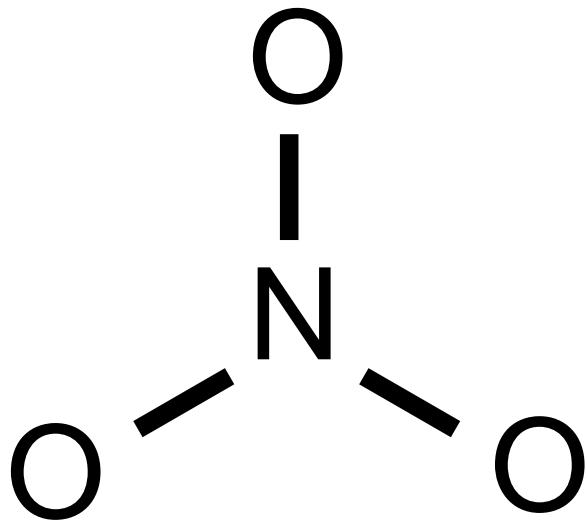


# Two-Color Resonant Four-Wave Mixing Spectroscopy of Jet Cooled $\text{NO}_3$ (II)

Masaru Fukushima and Takashi Ishiwata

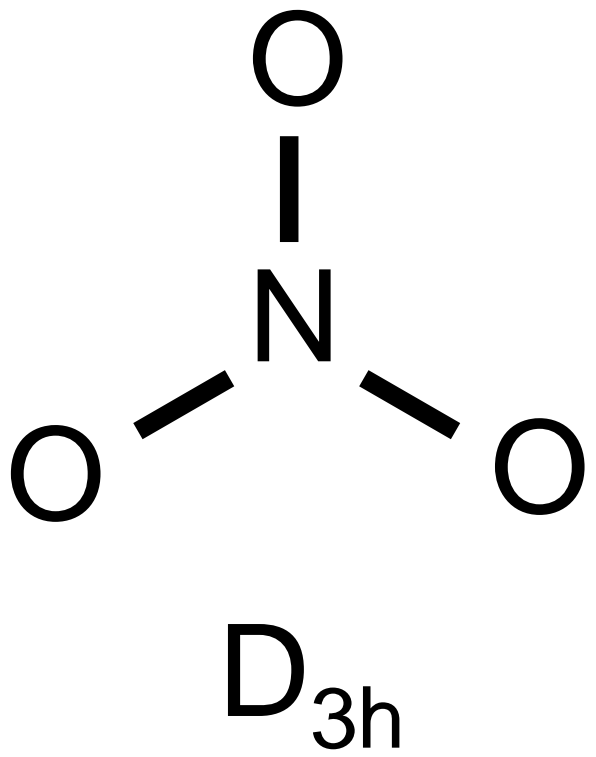
Graduate School of Information Sciences, Hiroshima City University,  
Hiroshima 731-3194, Japan

What is  $\text{NO}_3^-$  ?



$D_{3h}$

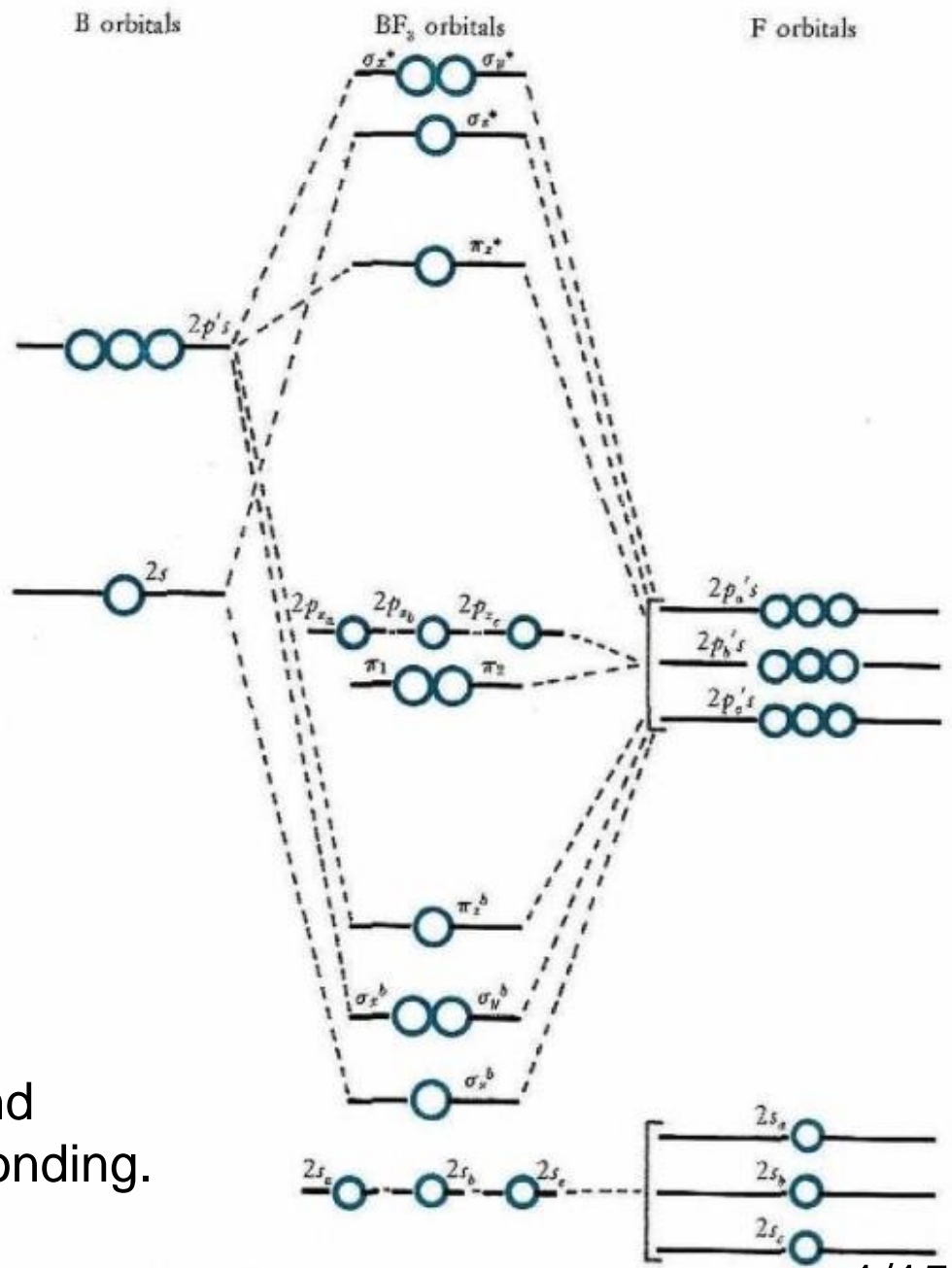
What is  $\text{NO}_3^-$  ?



Trigonal-Planar Molecules

$\text{BF}_3$

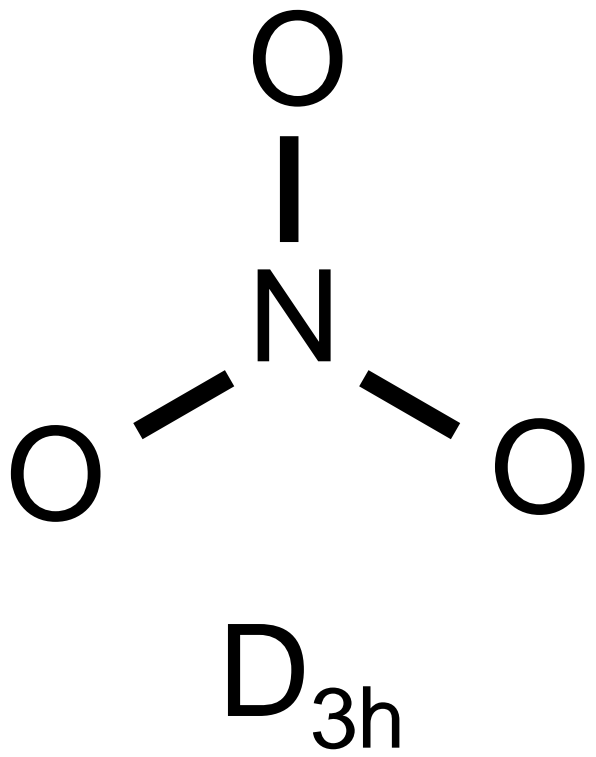
increasing energy



H. B. Gray,  
Electrons and  
Chemical Bonding.

Figure 4-7 Relative orbital energies in  $\text{BF}_3$ .

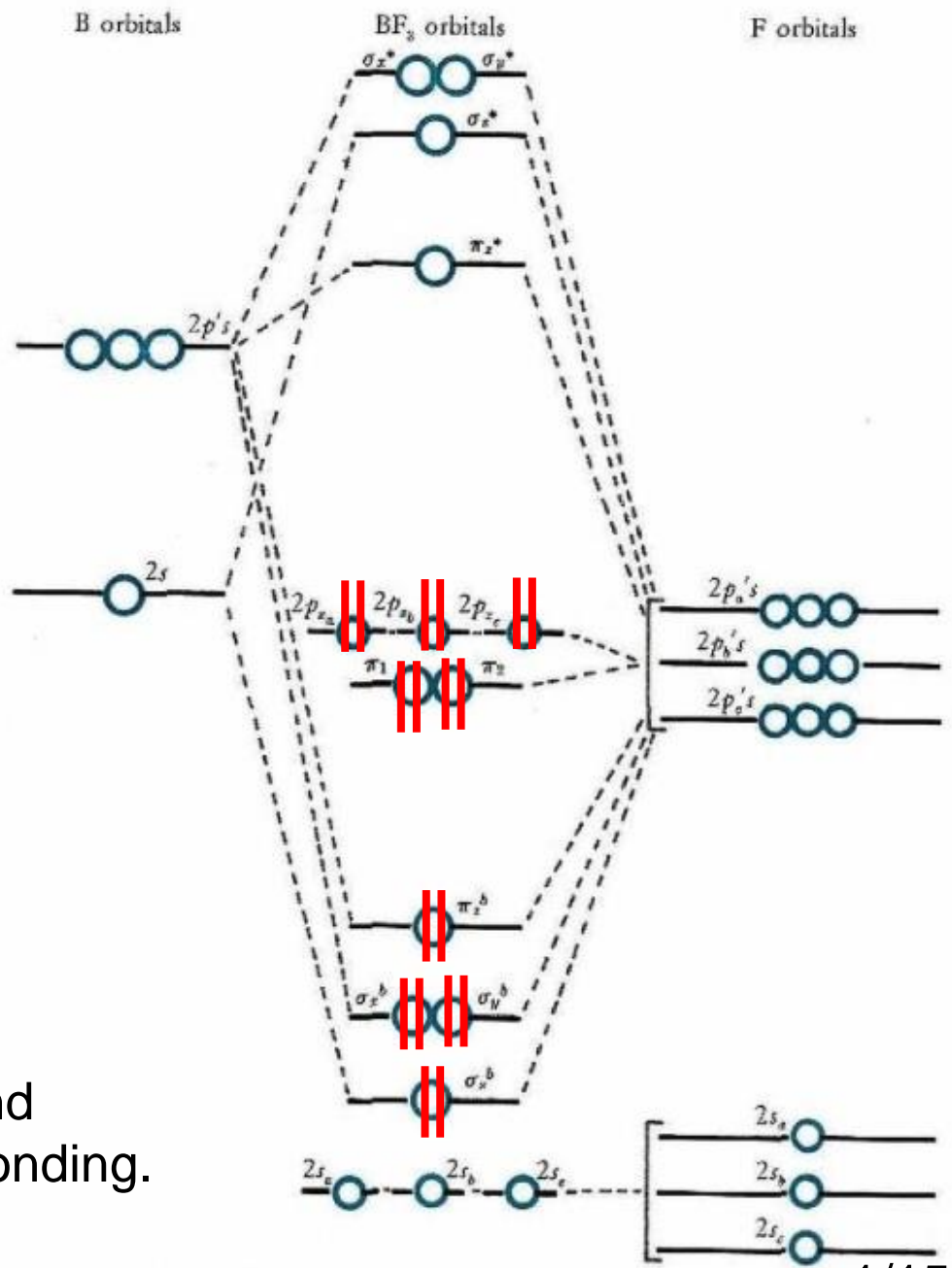
What is NO<sub>3</sub> ?



Trigonal-Planar Molecules

BF<sub>3</sub>

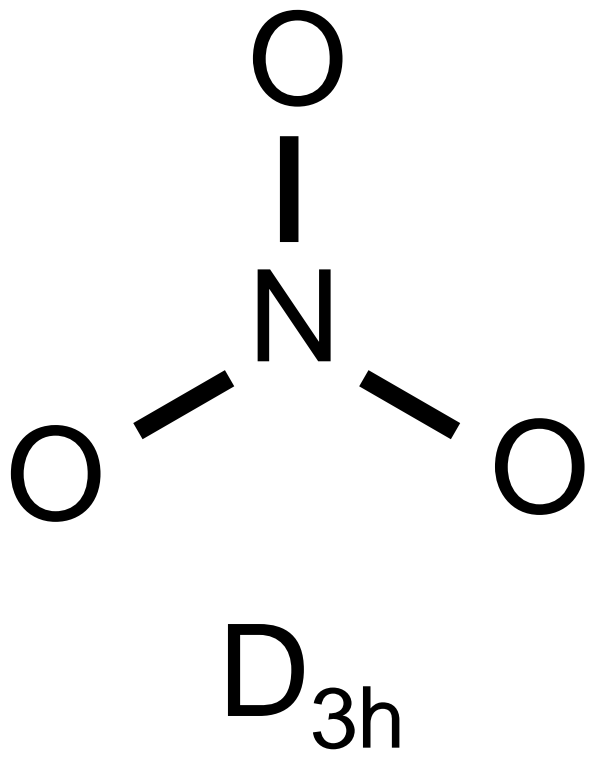
increasing energy



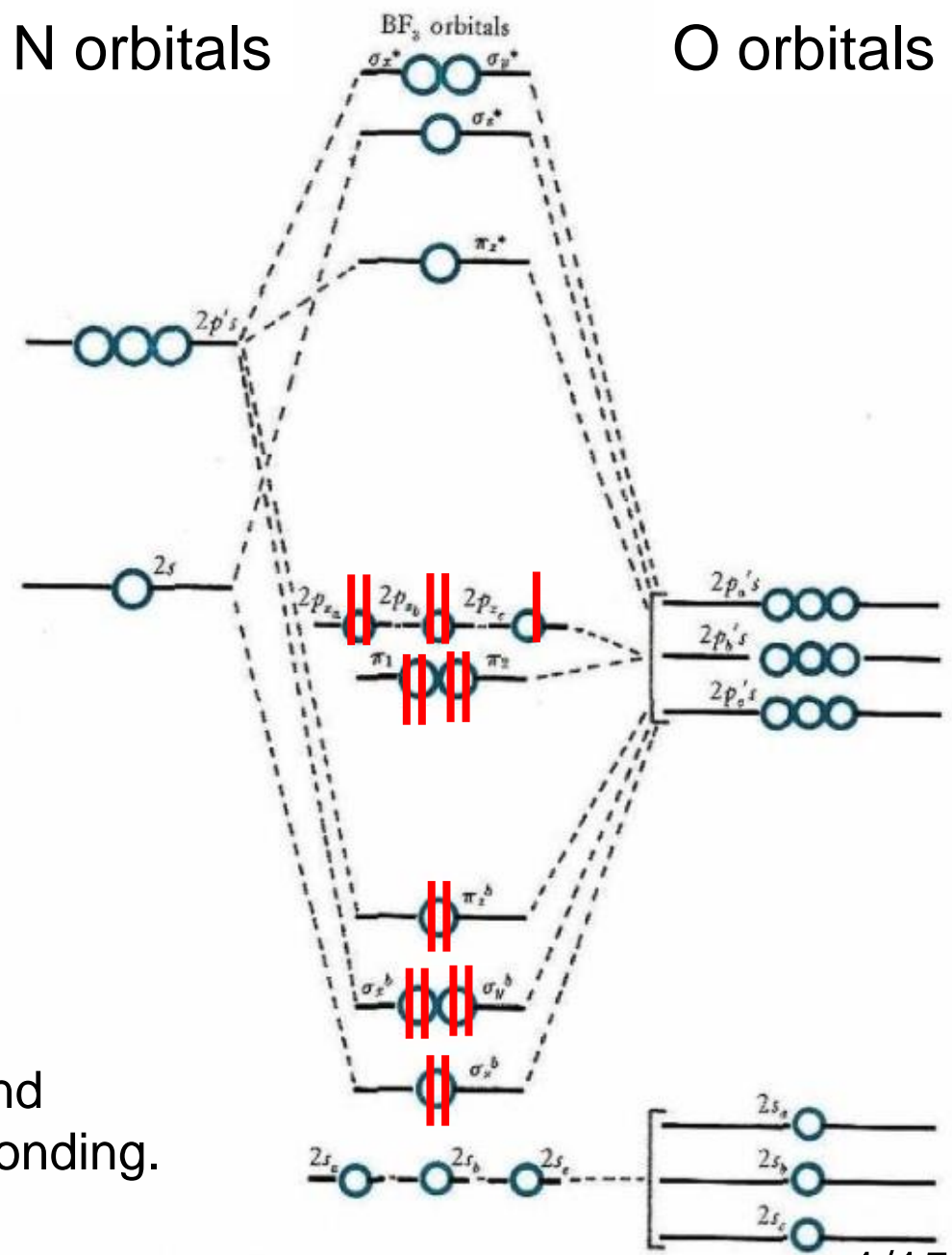
H. B. Gray,  
Electrons and  
Chemical Bonding.

Figure 4-7 Relative orbital energies in BF<sub>3</sub>.

# What is NO<sub>3</sub> ?



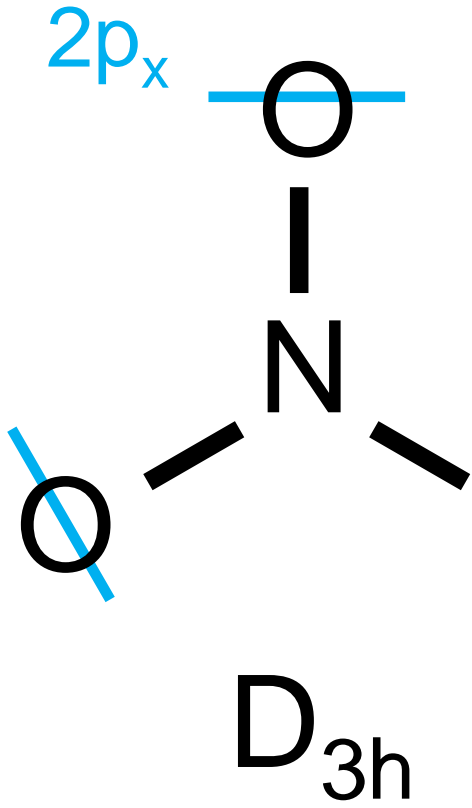
Trigonal-Planar Molecules



H. B. Gray,  
Electrons and  
Chemical Bonding.

Figure 4-7 Relative orbital energies in BF<sub>3</sub>.

# What is NO<sub>3</sub> ?



NO<sub>3</sub>

no contribution of N

H. B. Gray,  
Electrons and  
Chemical Bonding.

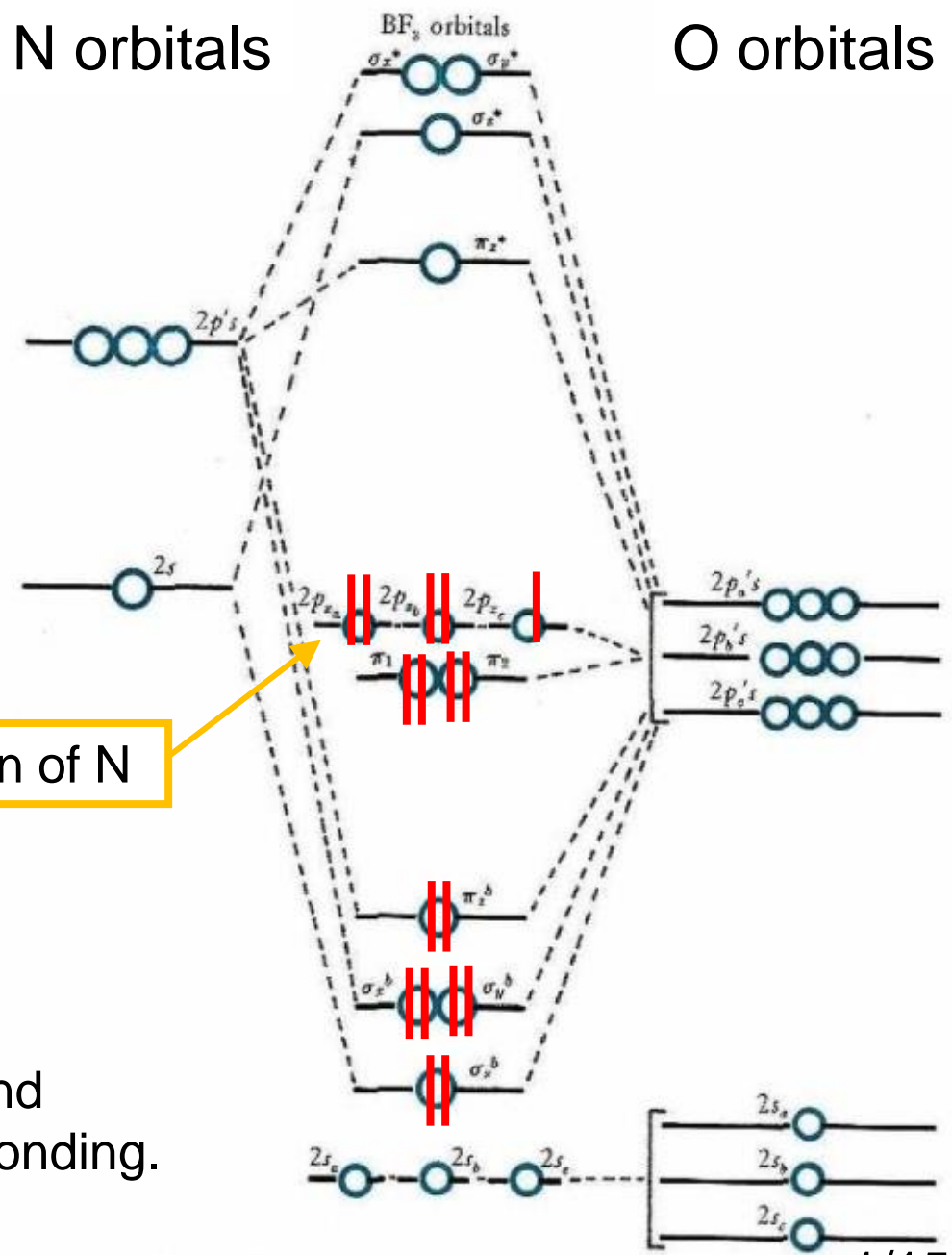
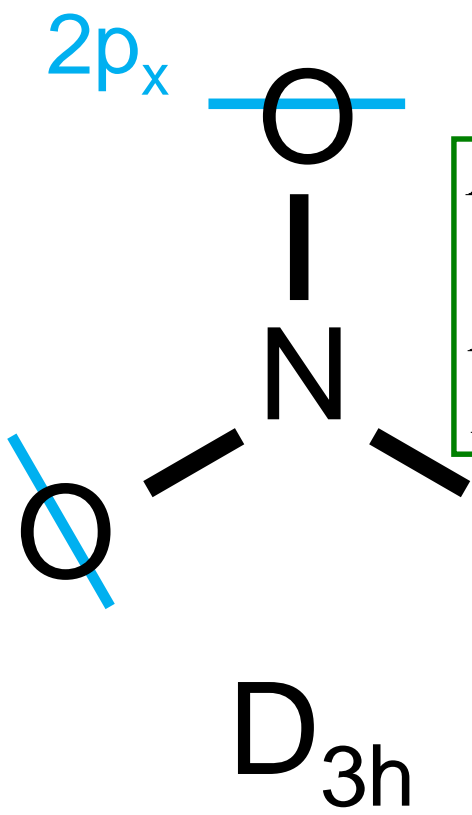


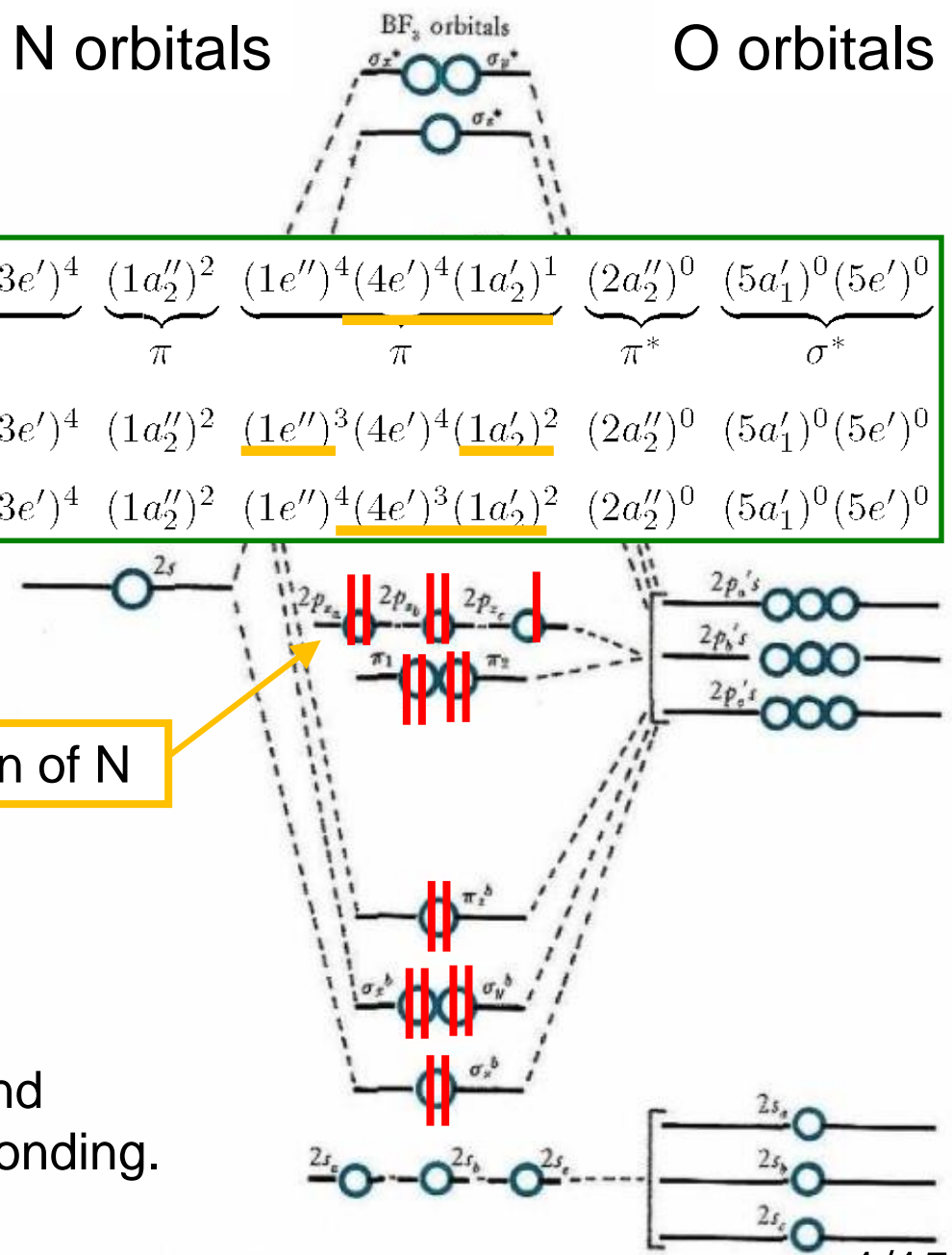
Figure 4-7 Relative orbital energies in BF<sub>3</sub>.

What is NO<sub>3</sub> ?



NO<sub>3</sub>

|                    |        |                    |              |                            |              |                    |
|--------------------|--------|--------------------|--------------|----------------------------|--------------|--------------------|
| $\tilde{X}^2 A'_2$ | (core) | $(4a'_1)^2(3e')^4$ | $(1a''_2)^2$ | $(1e'')^4(4e')^4(1a'_2)^1$ | $(2a''_2)^0$ | $(5a'_1)^0(5e')^0$ |
|                    |        | $\sigma$           | $\pi$        | $\pi$                      | $\pi^*$      | $\sigma^*$         |
| $\tilde{A}^2 E''$  | (core) | $(4a'_1)^2(3e')^4$ | $(1a''_2)^2$ | $(1e'')^3(4e')^4(1a'_2)^2$ | $(2a''_2)^0$ | $(5a'_1)^0(5e')^0$ |
| $\tilde{B}^2 E'$   | (core) | $(4a'_1)^2(3e')^4$ | $(1a''_2)^2$ | $(1e'')^4(4e')^3(1a'_2)^2$ | $(2a''_2)^0$ | $(5a'_1)^0(5e')^0$ |

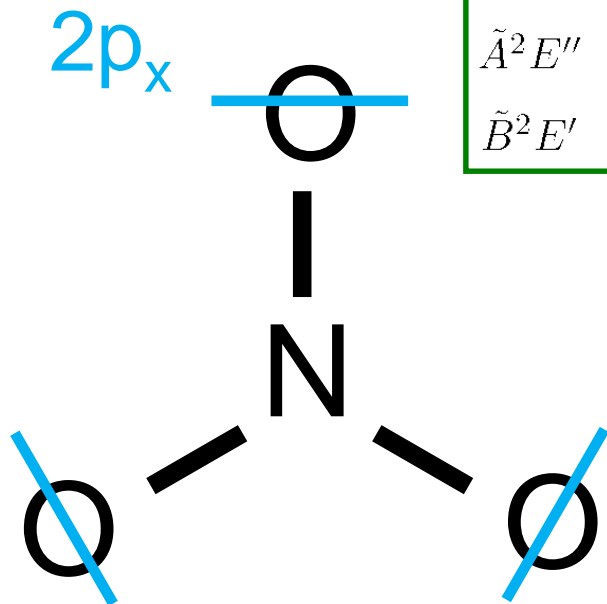


H. B. Gray,  
Electrons and  
Chemical Bonding.

Figure 4-7 Relative orbital energies in BF<sub>3</sub>. 1/15

# What is NO<sub>3</sub> ?

|                           |                    |             |                            |              |                    |
|---------------------------|--------------------|-------------|----------------------------|--------------|--------------------|
| $\tilde{X}^2 A'_2$ (core) | $(4a'_1)^2(3e')^4$ | $(1a'_2)^2$ | $(1e'')^4(4e')^4(1a'_2)^1$ | $(2a''_2)^0$ | $(5a'_1)^0(5e')^0$ |
|                           | $\sigma$           | $\pi$       | $\pi$                      | $\pi^*$      | $\sigma^*$         |
| $\tilde{A}^2 E''$ (core)  | $(4a'_1)^2(3e')^4$ | $(1a'_2)^2$ | $(1e'')^3(4e')^4(1a'_2)^2$ | $(2a''_2)^0$ | $(5a'_1)^0(5e')^0$ |
| $\tilde{B}^2 E'$ (core)   | $(4a'_1)^2(3e')^4$ | $(1a'_2)^2$ | $(1e'')^4(4e')^3(1a'_2)^2$ | $(2a''_2)^0$ | $(5a'_1)^0(5e')^0$ |



$D_{3h}$

Eisfeld and Morokuma,  
J. Chem. Phys. 114,  
9430 (2001).

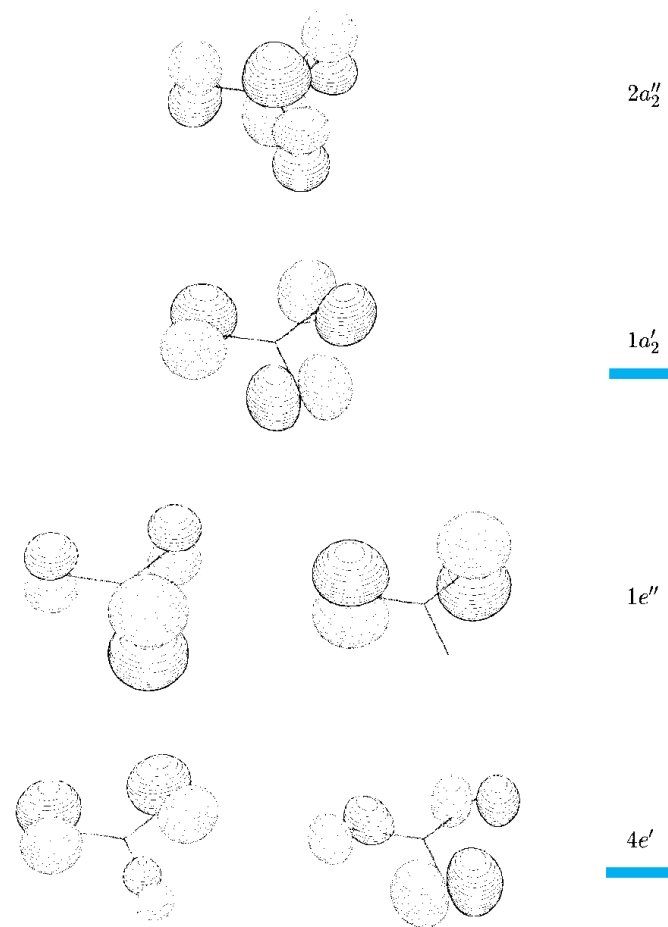
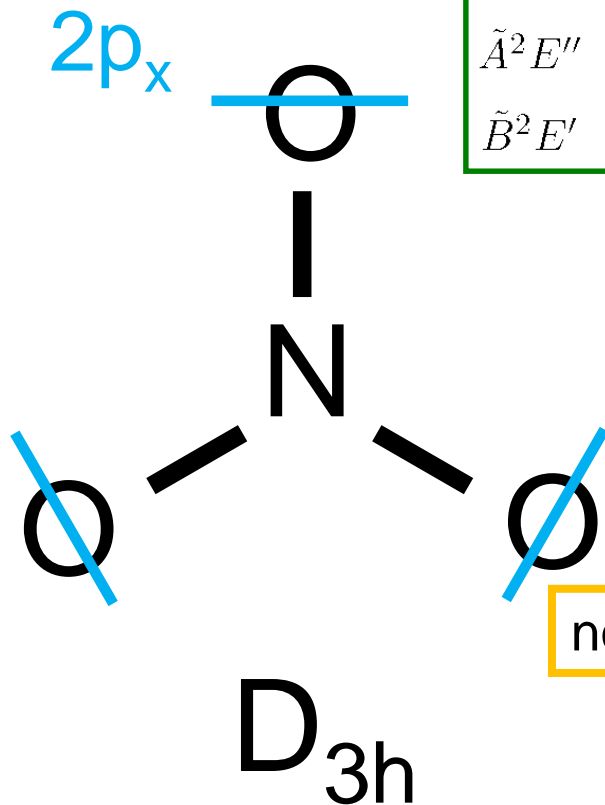


FIG. 1. Selected molecular orbitals of NO<sub>3</sub> as calculated by CASSCF at  $D_{3h}$  geometry with a N–O distance of 1.240 Å.

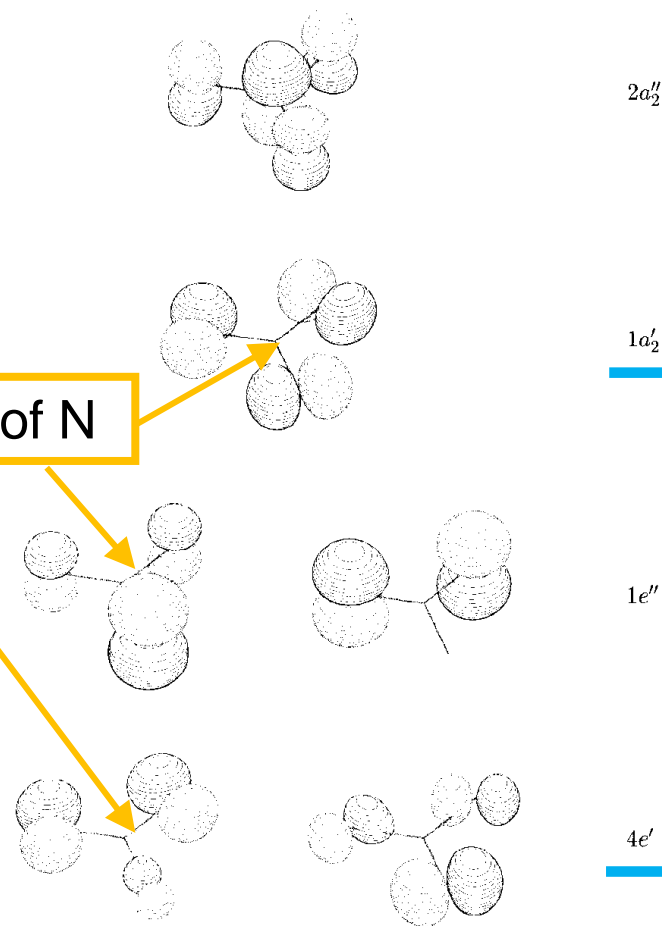


# What is NO<sub>3</sub> ?

|                           |                    |              |                            |              |                    |
|---------------------------|--------------------|--------------|----------------------------|--------------|--------------------|
| $\tilde{X}^2 A'_2$ (core) | $(4a'_1)^2(3e')^4$ | $(1a''_2)^2$ | $(1e'')^4(4e')^4(1a'_2)^1$ | $(2a''_2)^0$ | $(5a'_1)^0(5e')^0$ |
|                           | $\sigma$           | $\pi$        | $\pi$                      | $\pi^*$      | $\sigma^*$         |
| $\tilde{A}^2 E''$ (core)  | $(4a'_1)^2(3e')^4$ | $(1a''_2)^2$ | $(1e'')^3(4e')^4(1a'_2)^2$ | $(2a''_2)^0$ | $(5a'_1)^0(5e')^0$ |
| $\tilde{B}^2 E'$ (core)   | $(4a'_1)^2(3e')^4$ | $(1a''_2)^2$ | $(1e'')^4(4e')^3(1a'_2)^2$ | $(2a''_2)^0$ | $(5a'_1)^0(5e')^0$ |



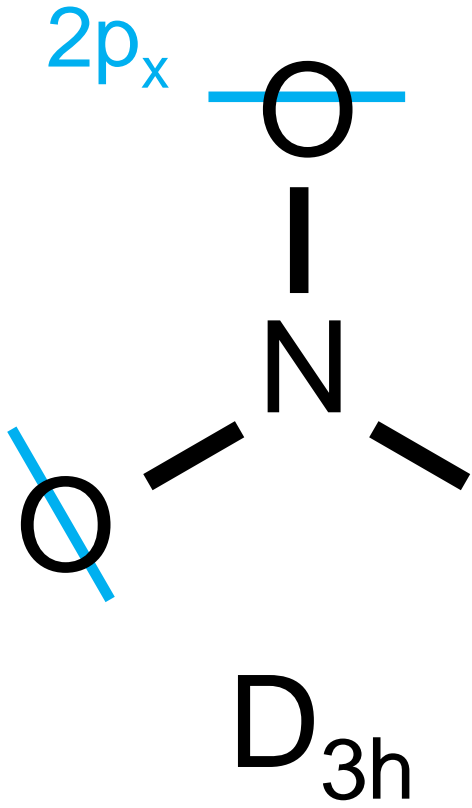
no contribution of N



Eisfeld and Morokuma,  
J. Chem. Phys. 114,  
9430 (2001).

FIG. 1. Selected molecular orbitals of NO<sub>3</sub> as calculated by CASSCF at  $D_{3h}$  geometry with a N–O distance of 1.240 Å.

# What is NO<sub>3</sub> ?



NO<sub>3</sub>

no contribution of N

H. B. Gray,  
Electrons and  
Chemical Bonding.

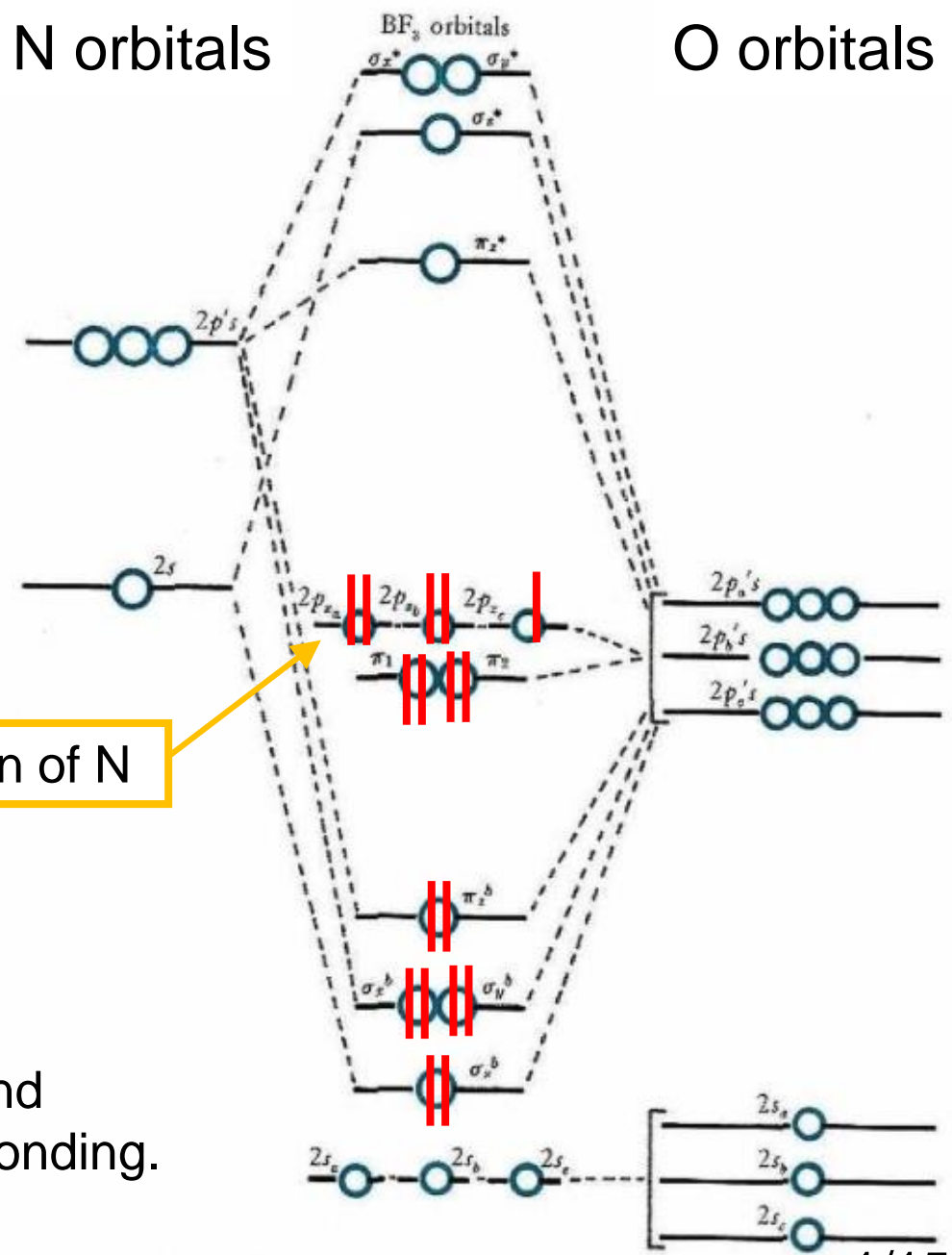
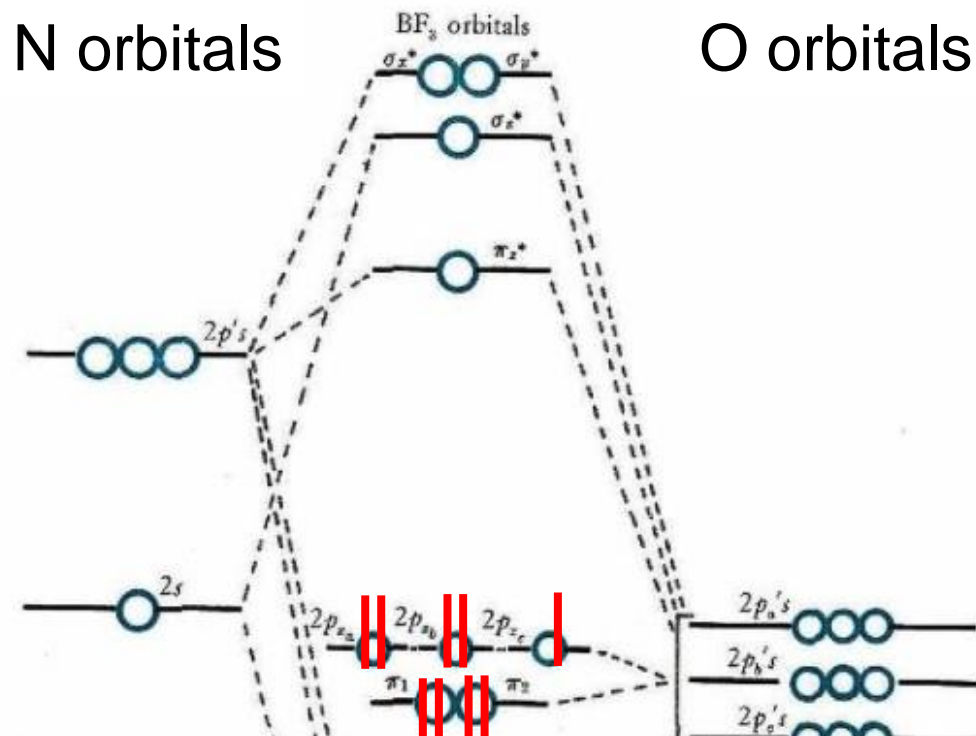
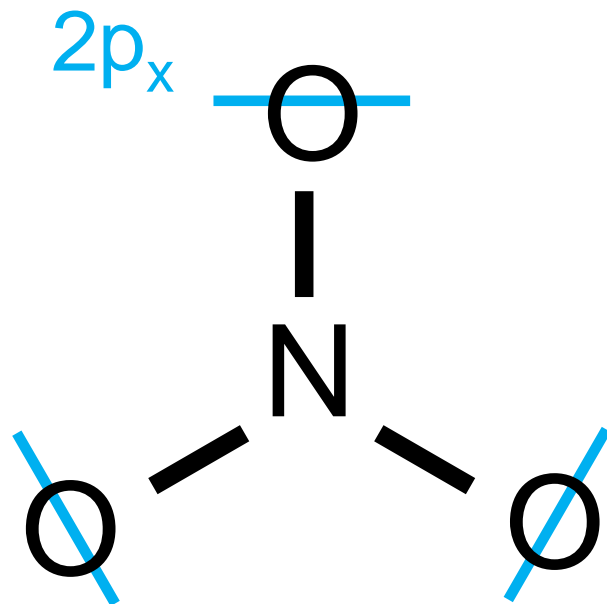


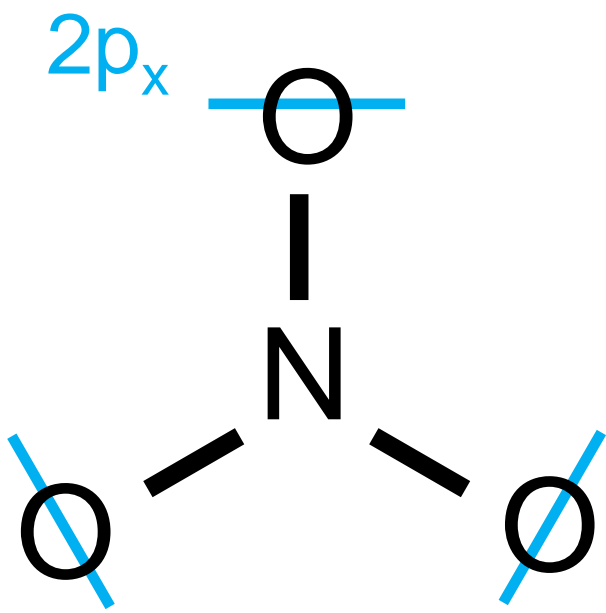
Figure 4-7 Relative orbital energies in BF<sub>3</sub>. 1/15

# What is $\text{NO}_3^-$ ?

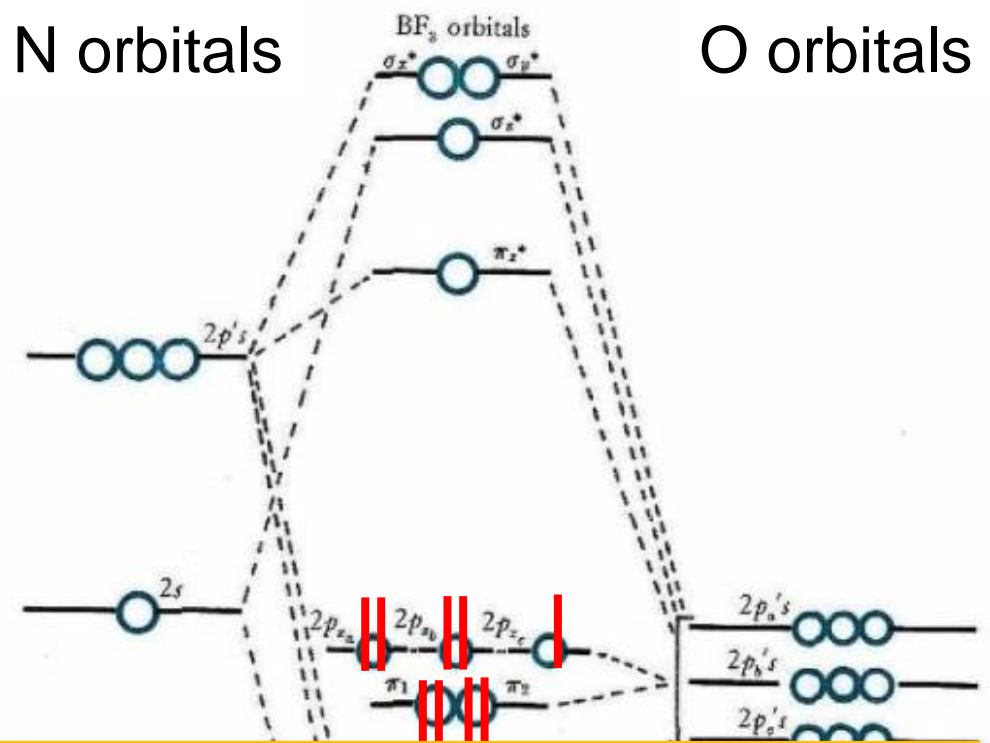


1. SOMO's (singly occupied molecular orbital) have little contribution of N.
2. The un-paired electron is localized on O's.
  - ⇒ SOMO just consists of 2p of O, and has no chemical bond.
  - ⇒ The un-paired electron easily follows the motion of O's.
  - ⇒ Vibronic couplings in the states are strong.

What is  $\text{NO}_3^-$  ?

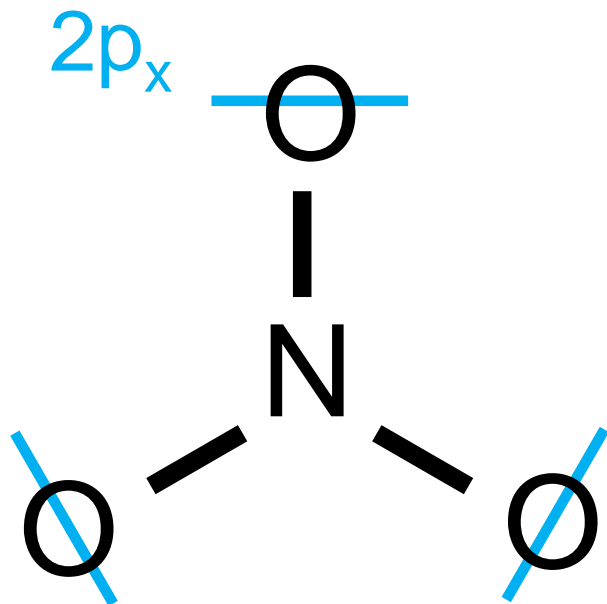


$\text{NO}_3^-$

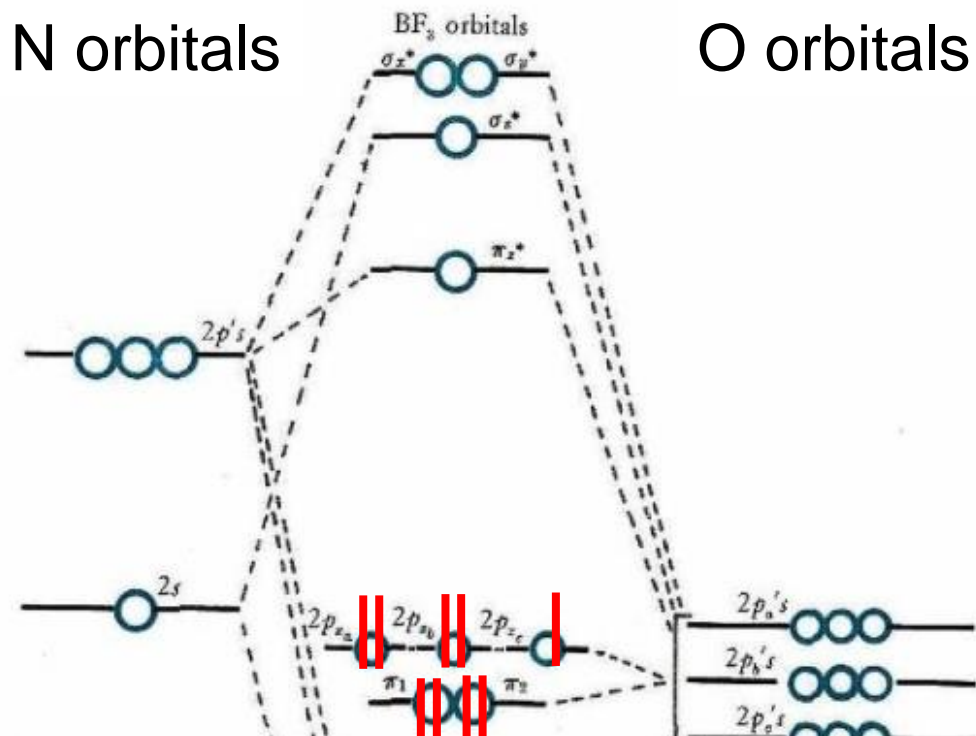


1. SOMO's (singly occupied molecular orbital) have little contribution of N.
2. The un-paired electron is localized on O's.
  - ⇒ SOMO just consists of 2p of O, and has no chemical bond.
  - ⇒ The un-paired electron easily follows the motion of O's.
  - ⇒ Vibronic couplings in the states are strong.

# What is $\text{NO}_3^-$ ?

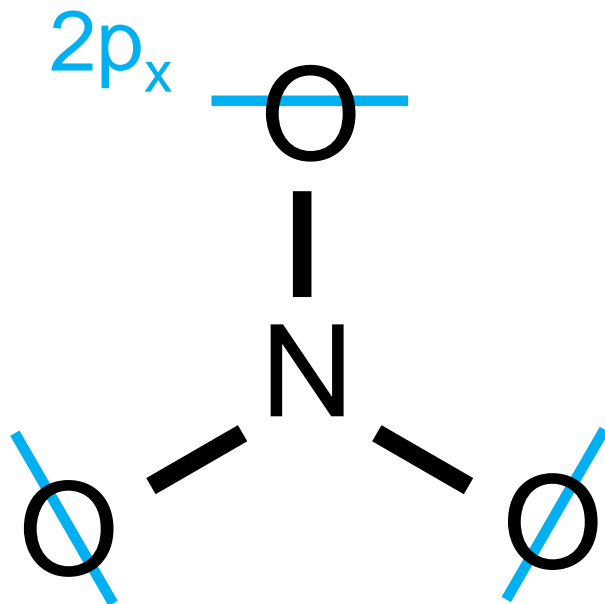


$\text{NO}_3^-$

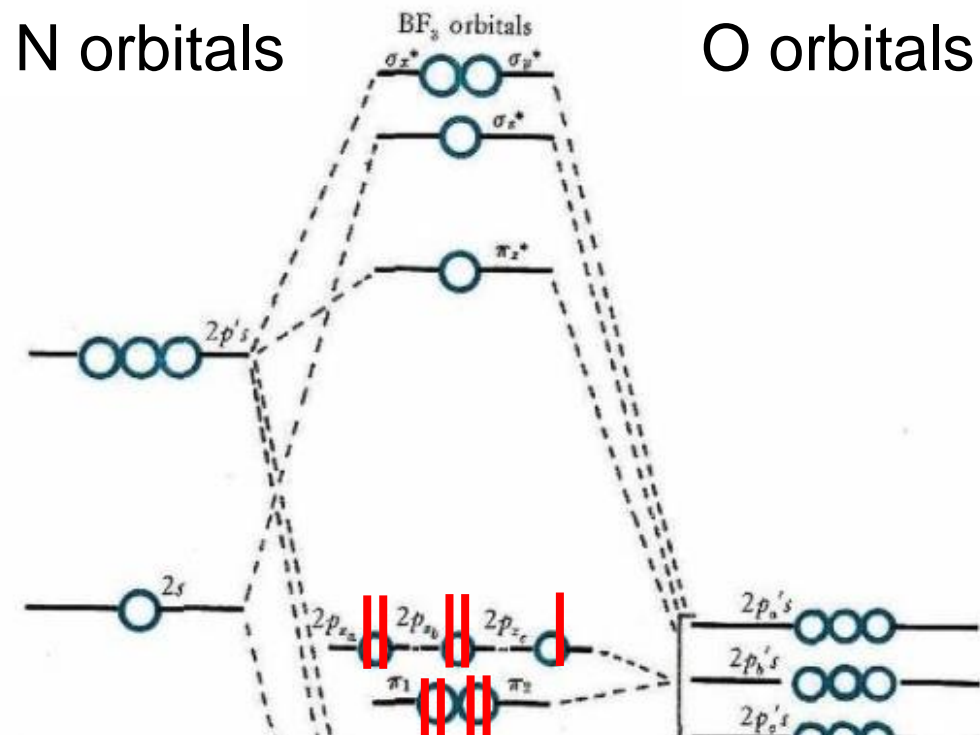


1. SOMO's (singly occupied molecular orbital) have little contribution of N.
2. The un-paired electron is localized on O's.
  - ⇒ SOMO just consists of 2p of O, and has no chemical bond.
  - ⇒ The un-paired electron easily follows the motion of O's.
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# What is $\text{NO}_3^-$ ?



$\text{NO}_3^-$

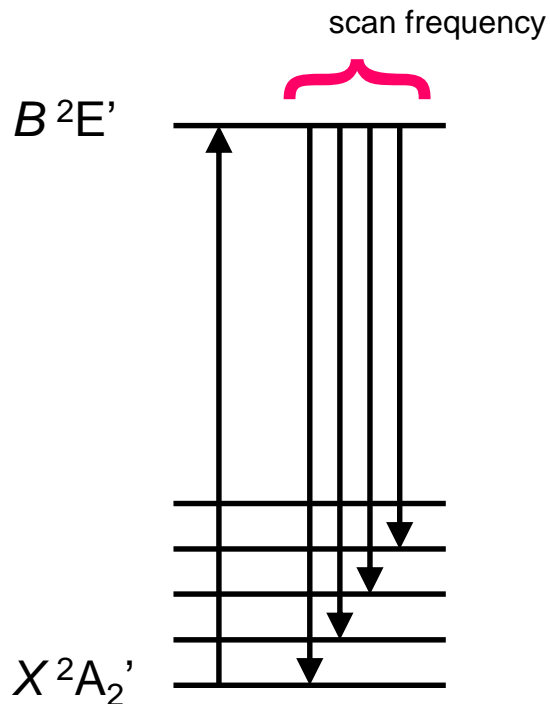


1. SOMO's (singly occupied molecular orbital) have little contribution of N.
2. The un-paired electron is localized on O's.
  - ⇒ SOMO just consists of 2p of O, and has no chemical bond.
  - ⇒ The un-paired electron easily follows the motion of O's.
  - ⇒ Vibronic couplings in the states are strong.

# Why 4WM ?

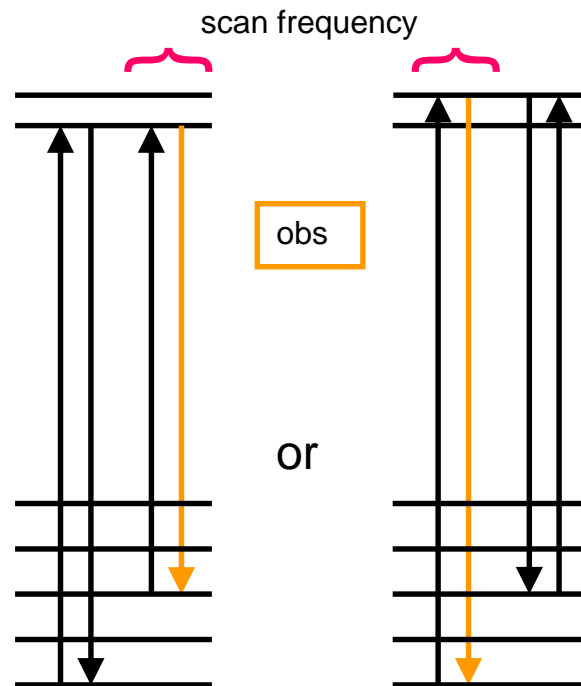
## Spectral resolution

Dispersed fluorescence spec.



resolution : monochromator  
( vibrationally resolved )

Four-wave mixing spec.  
( 4WM spec. )

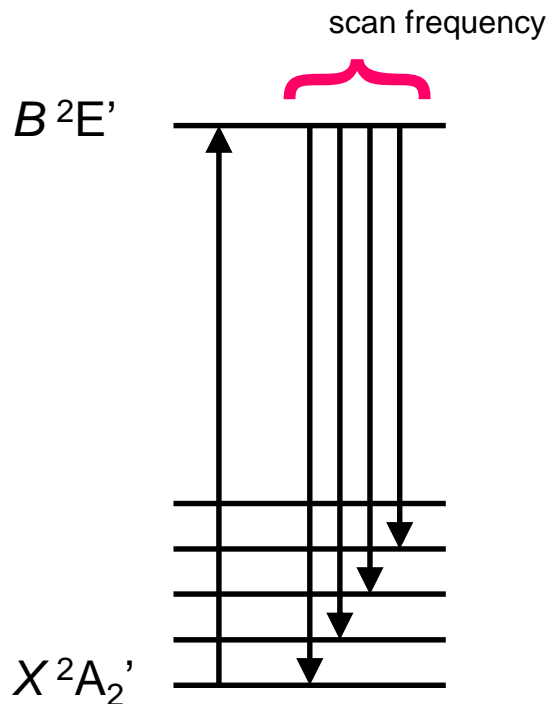


resolution : excitation laser  
( rotationally resolved )

# Why 4WM ?

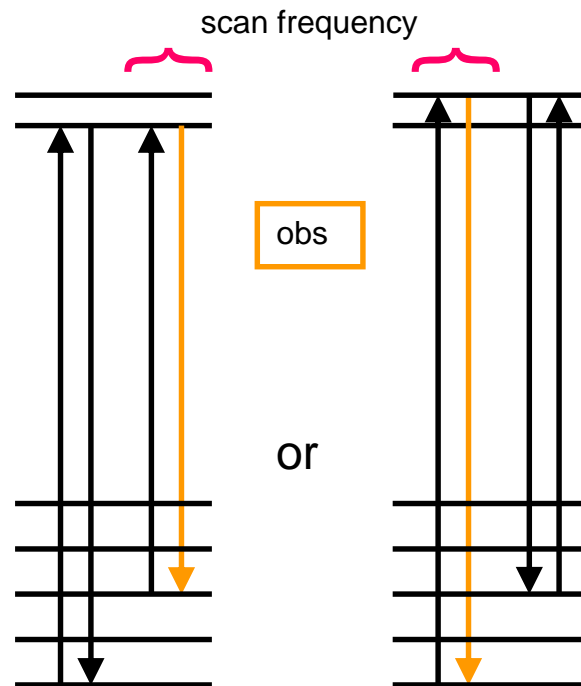
We use 4WM to measure rotationally resolved DF spectra.

Dispersed fluorescence spec.



resolution : monochromator  
( vibrationally resolved )

Four-wave mixing spec.  
( 4WM spec. )

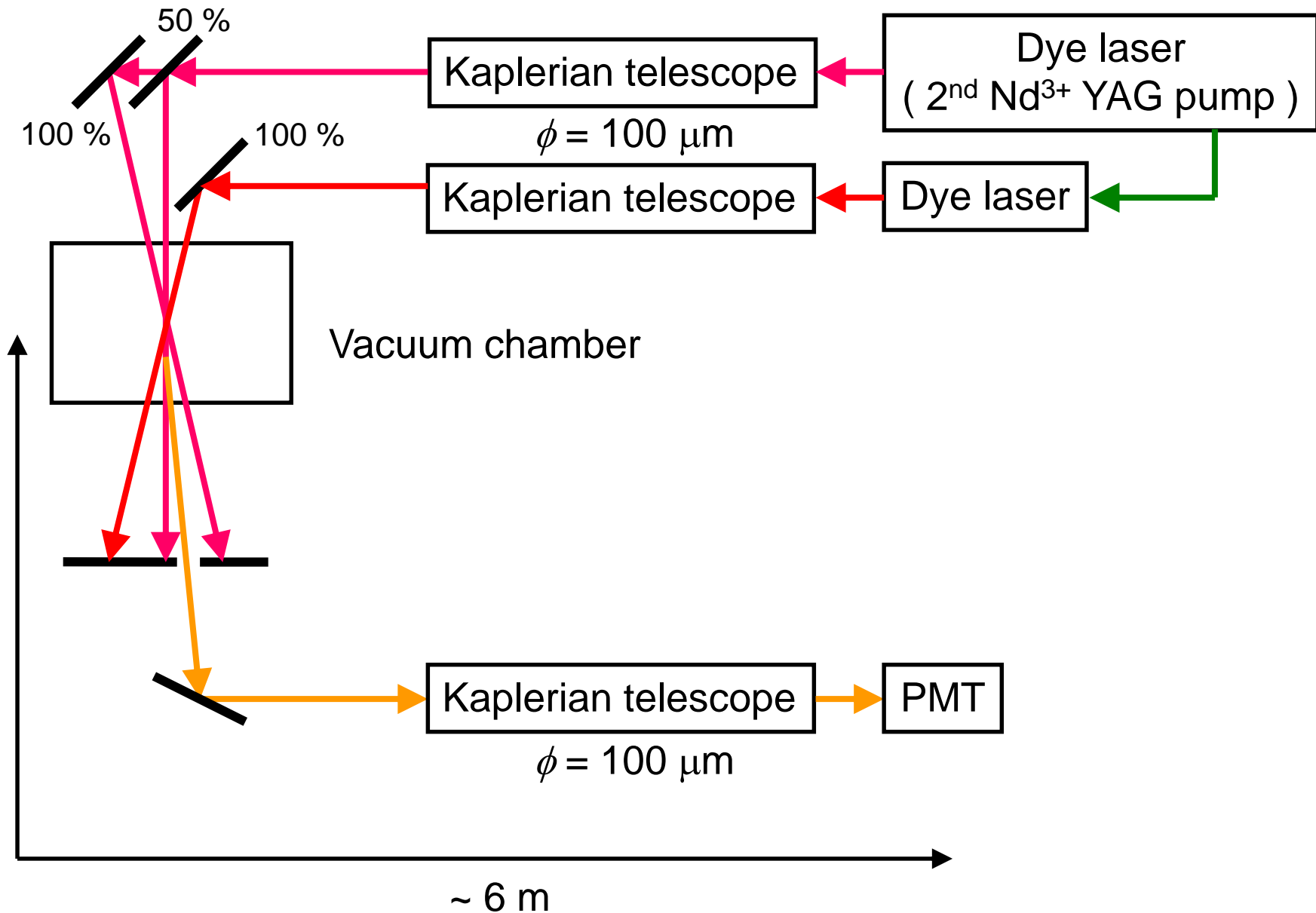


resolution : excitation laser  
( rotationally resolved )

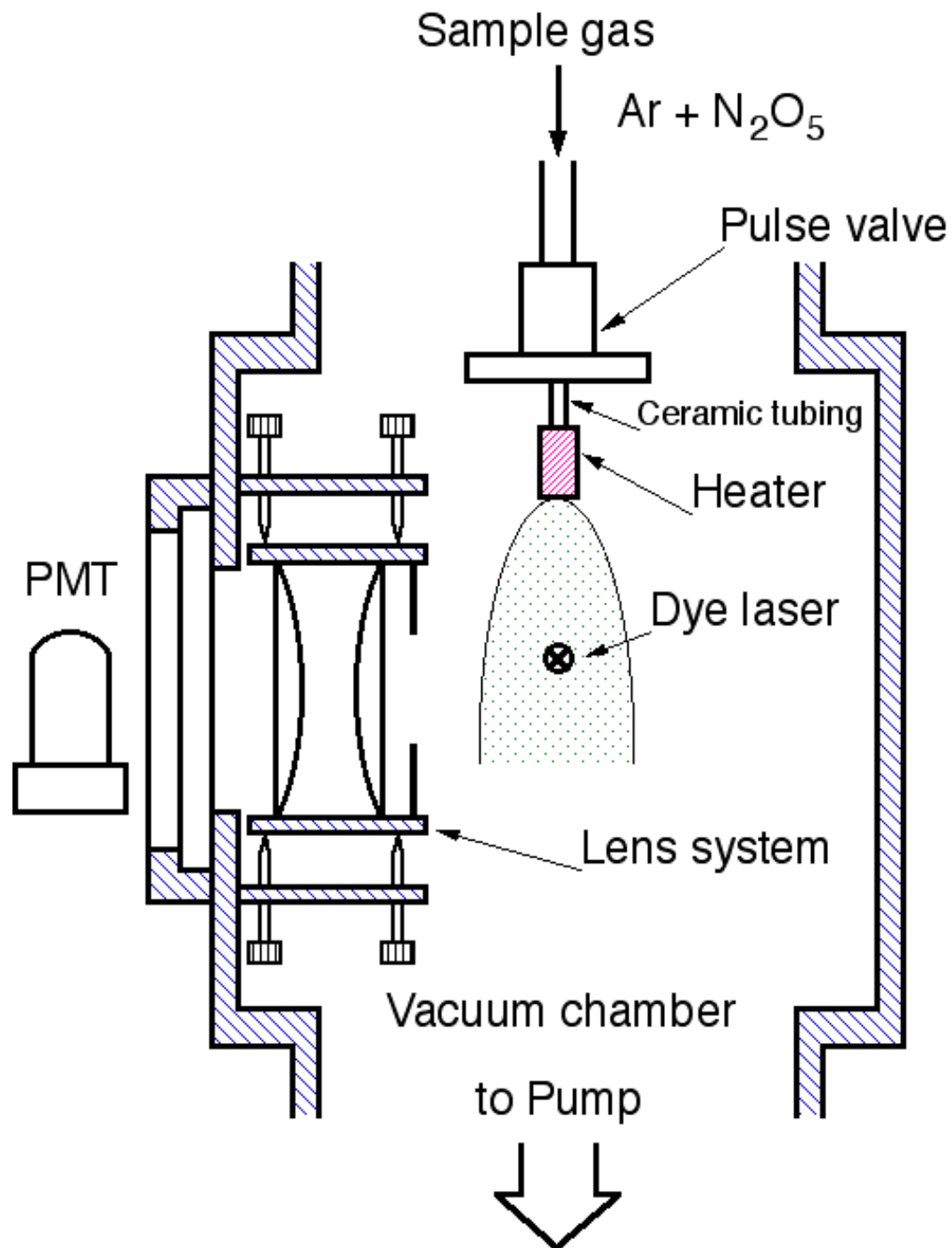


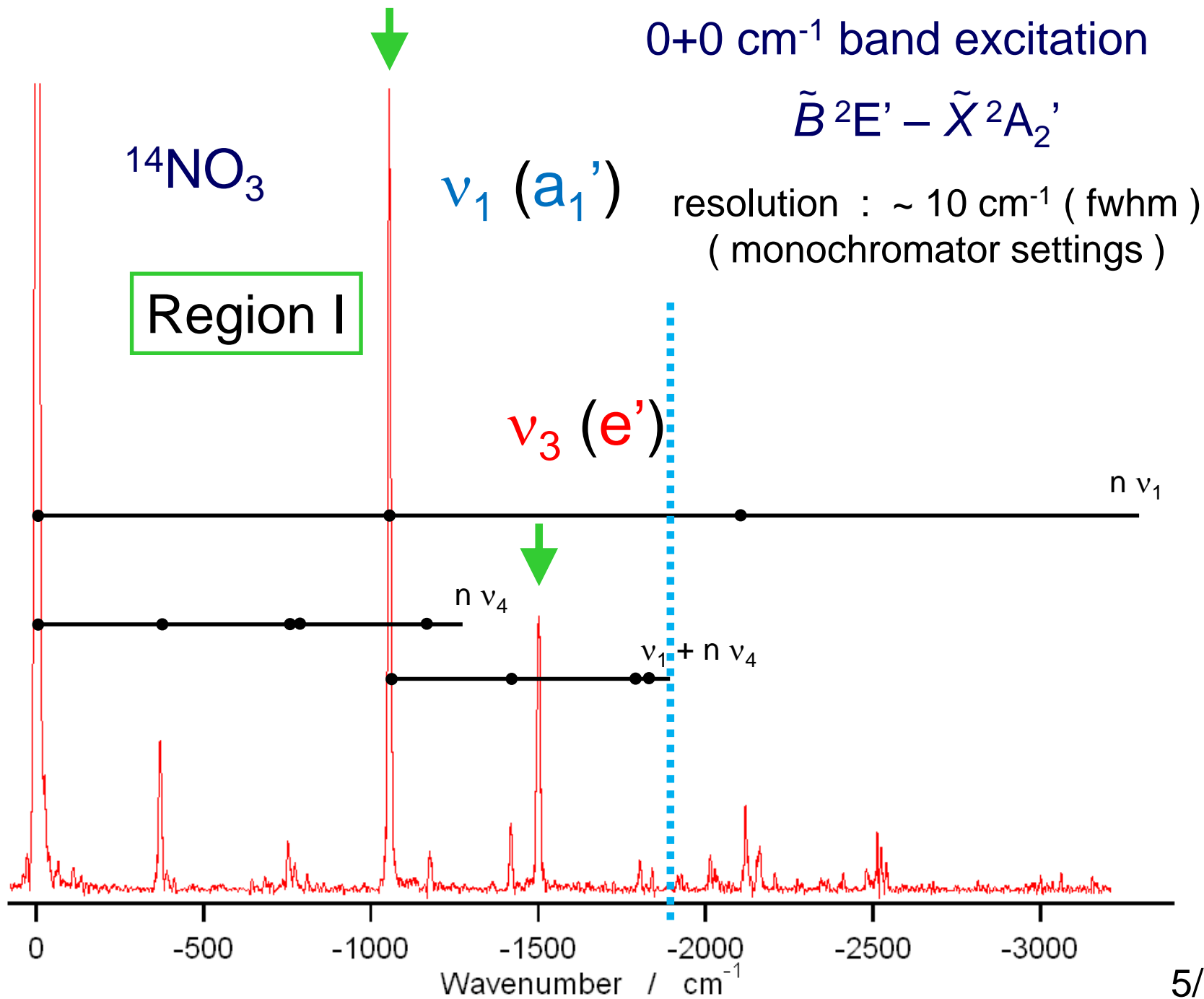
# Experiment

for 2C-R4WM

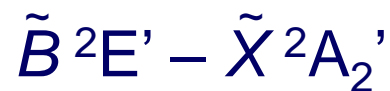


# Experiment





0+0 cm<sup>-1</sup> band excitation



<sup>14</sup>NO<sub>3</sub>

$\nu_1$  ( $a_1'$ )  
(IR forbidden)

resolution : ~ 10 cm<sup>-1</sup> ( fwhm )  
( monochromator settings )

Region I

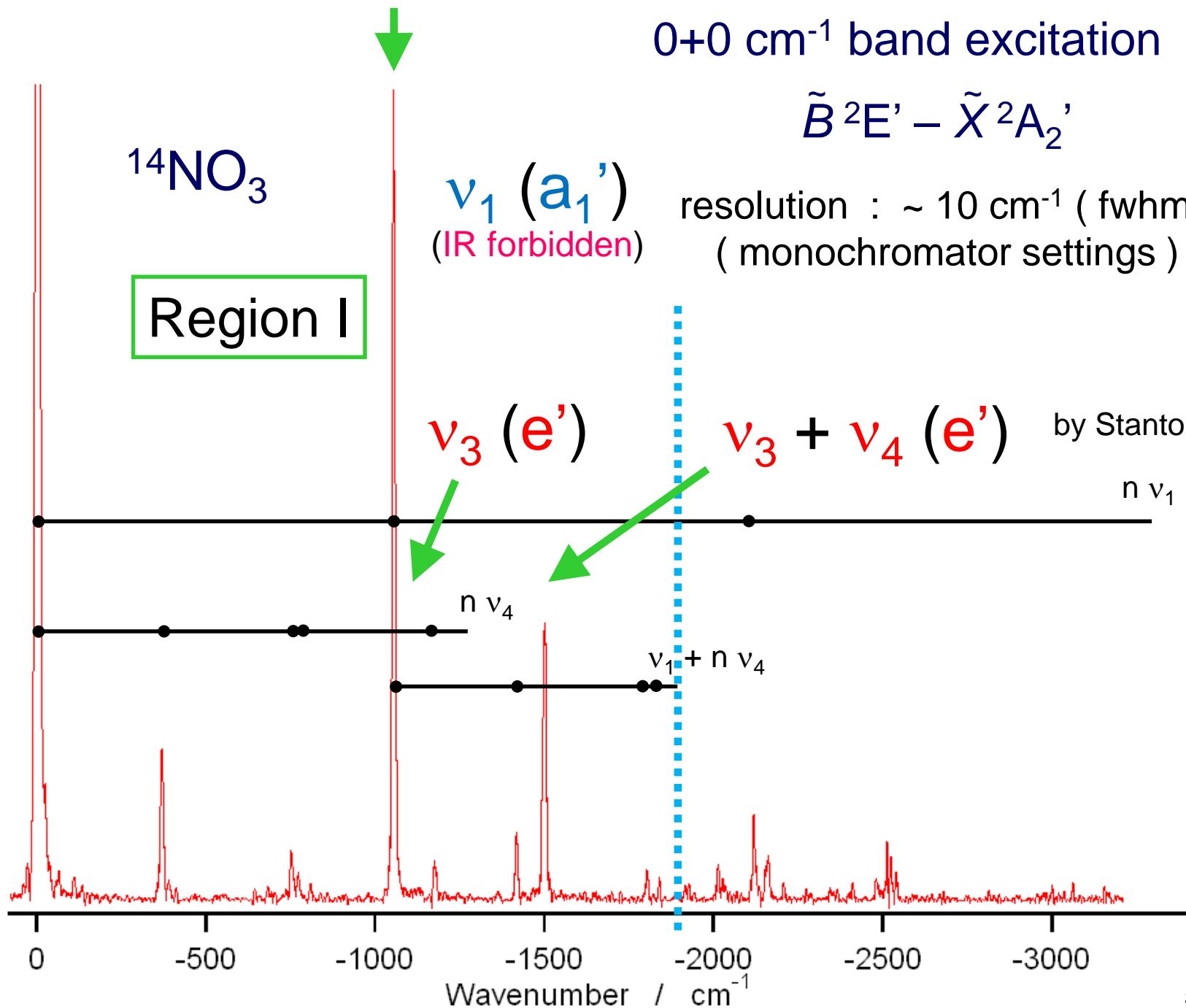
$\nu_3$  ( $e'$ )

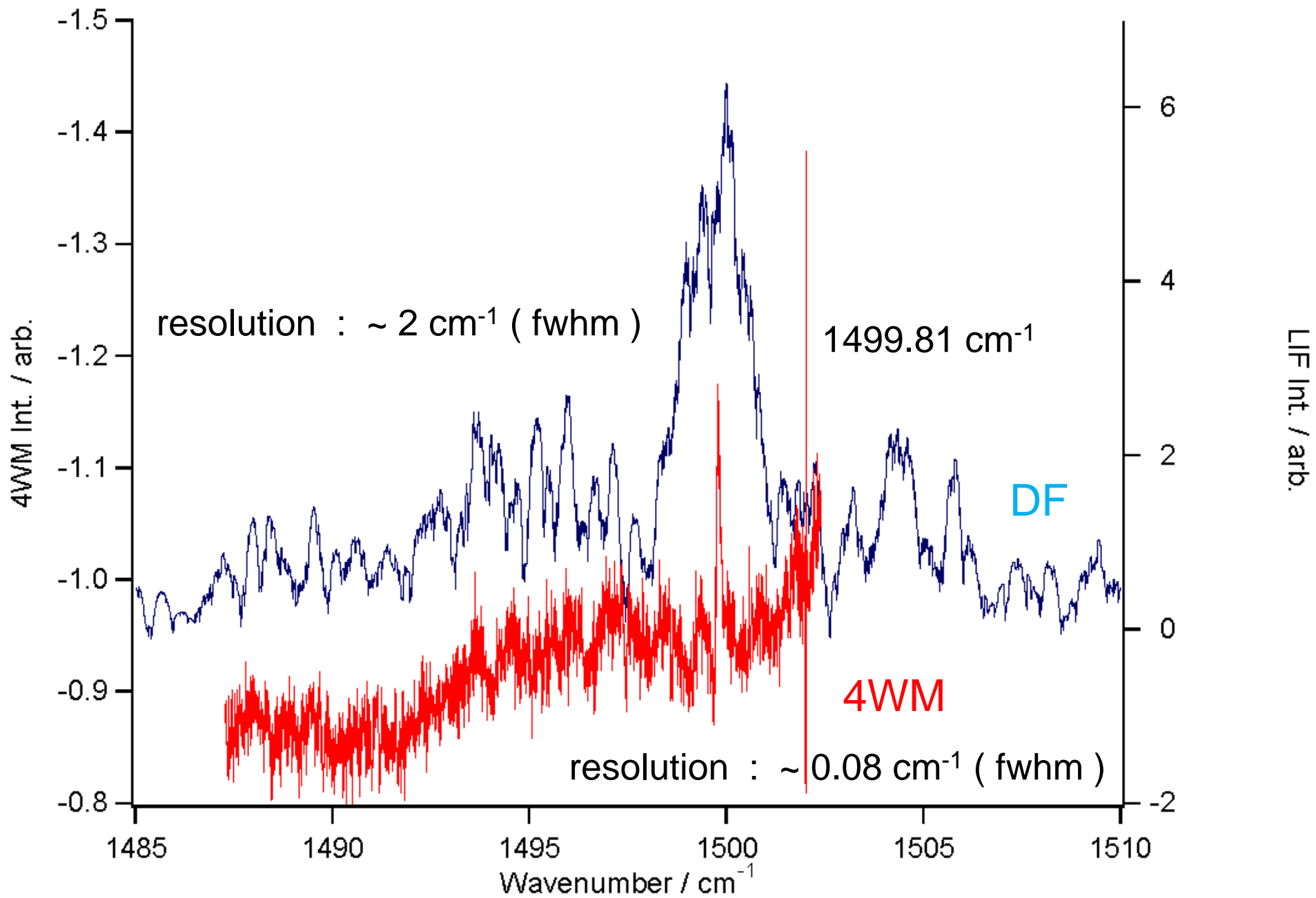
$\nu_3 + \nu_4$  ( $e'$ ) by Stanton

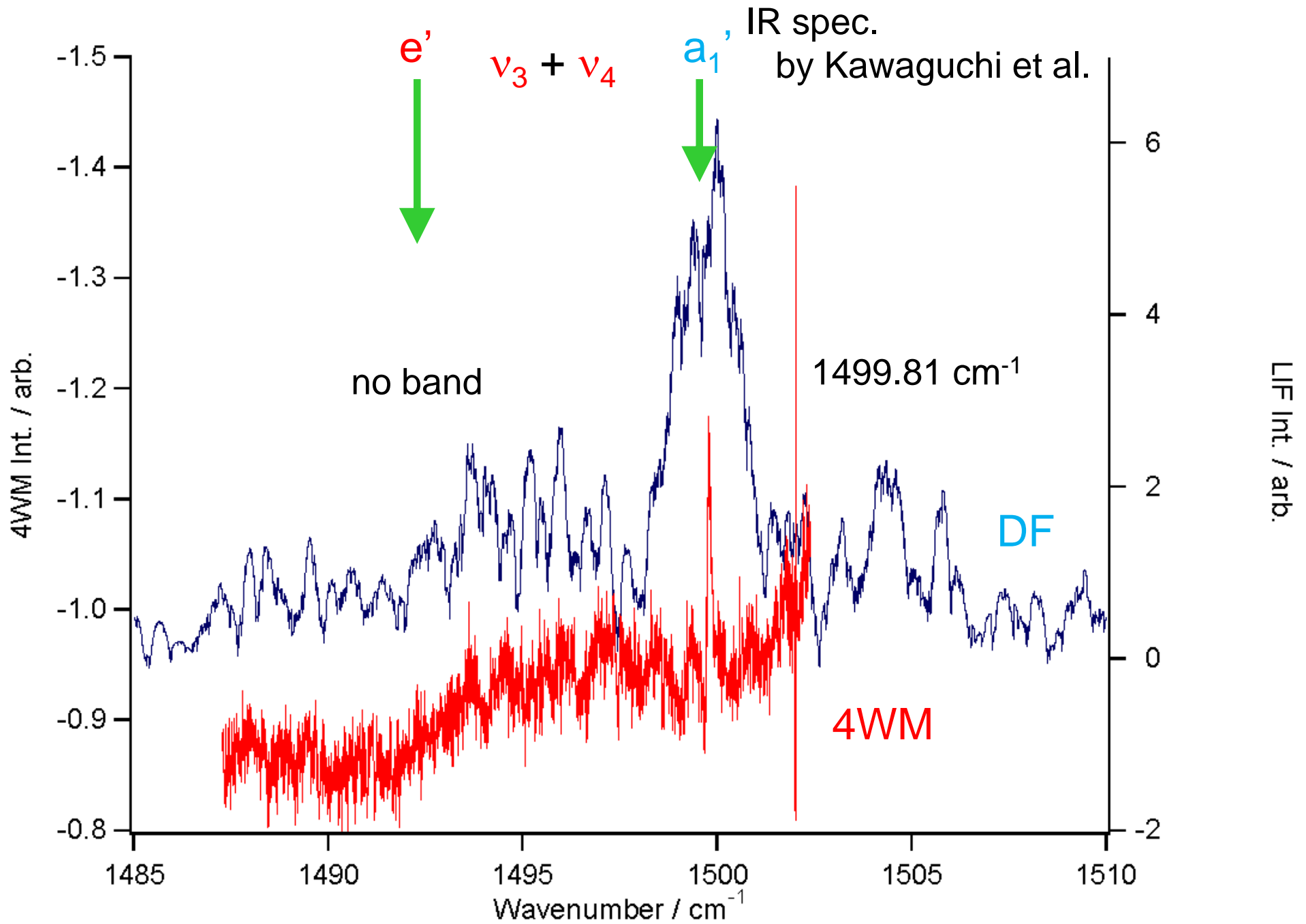
$n \nu_1$

$n \nu_4$

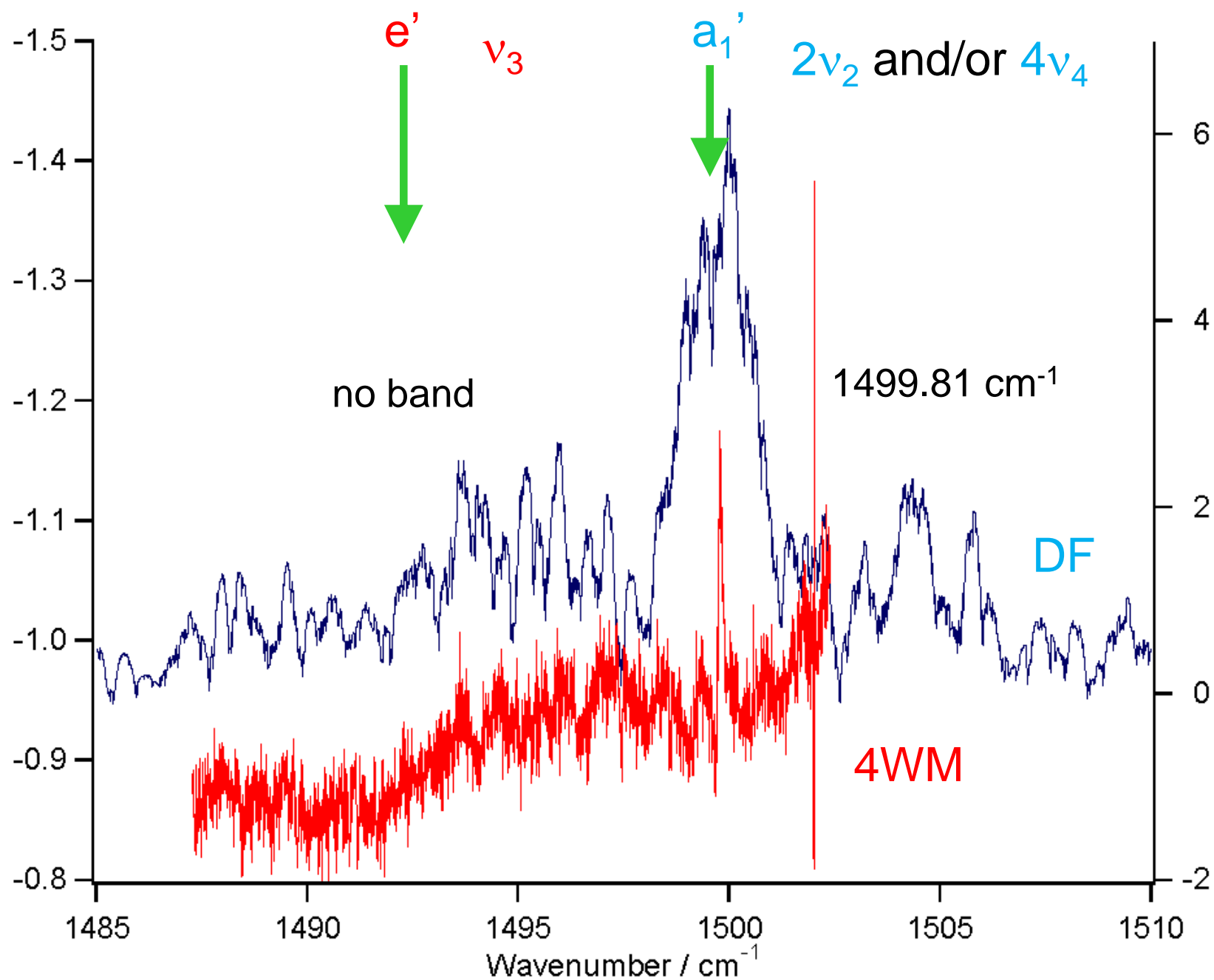
$\nu_1 + n \nu_4$



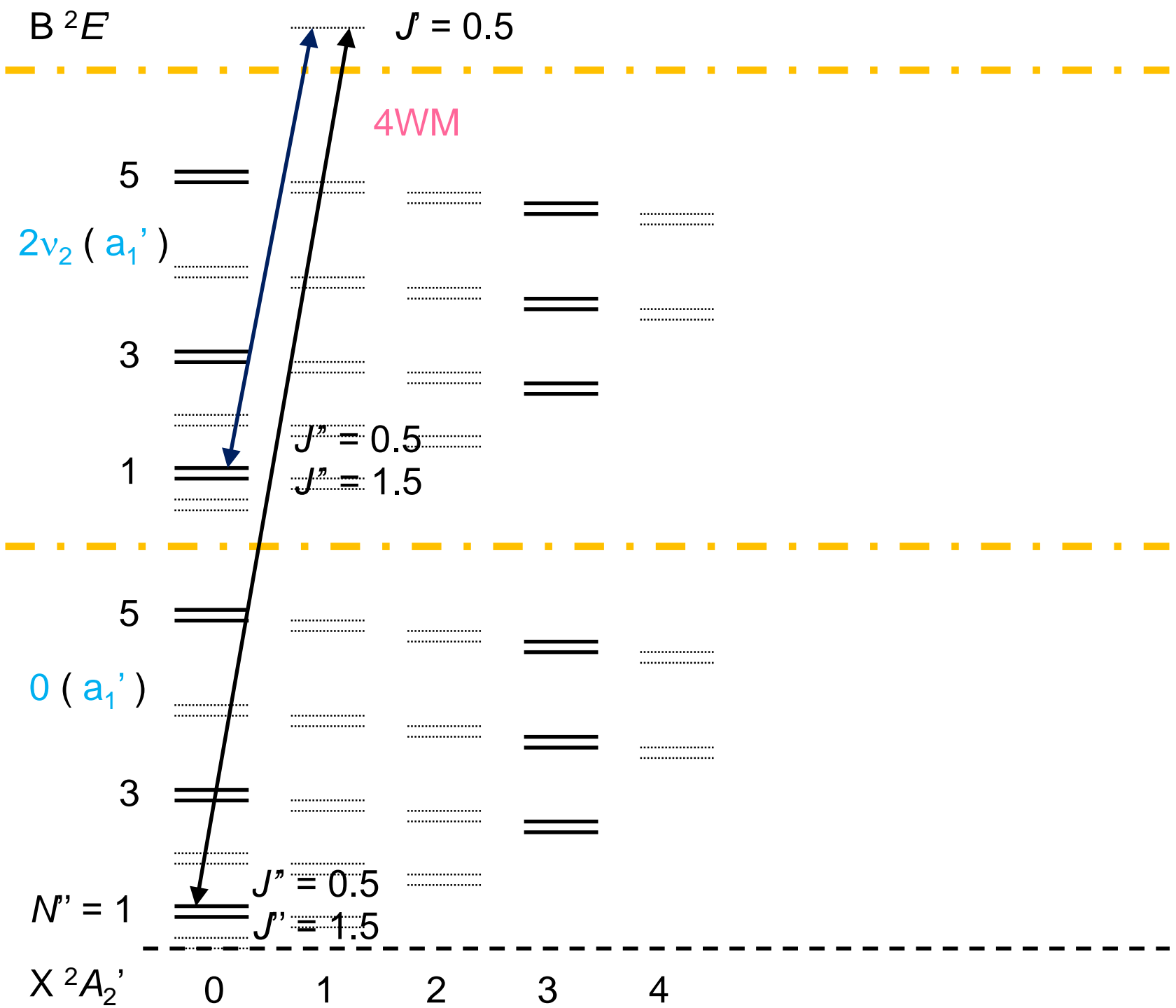




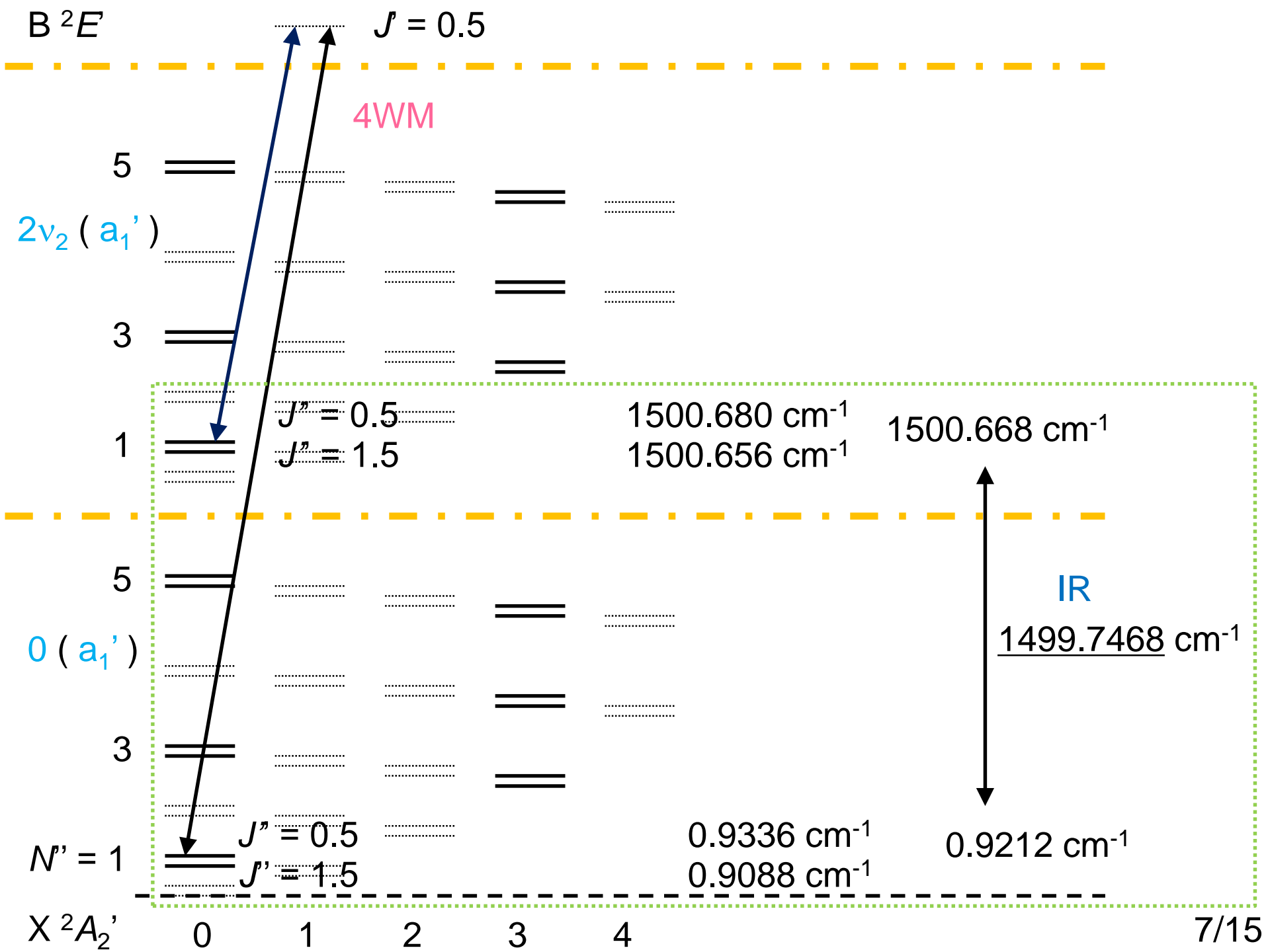
4WM Int. / arb.

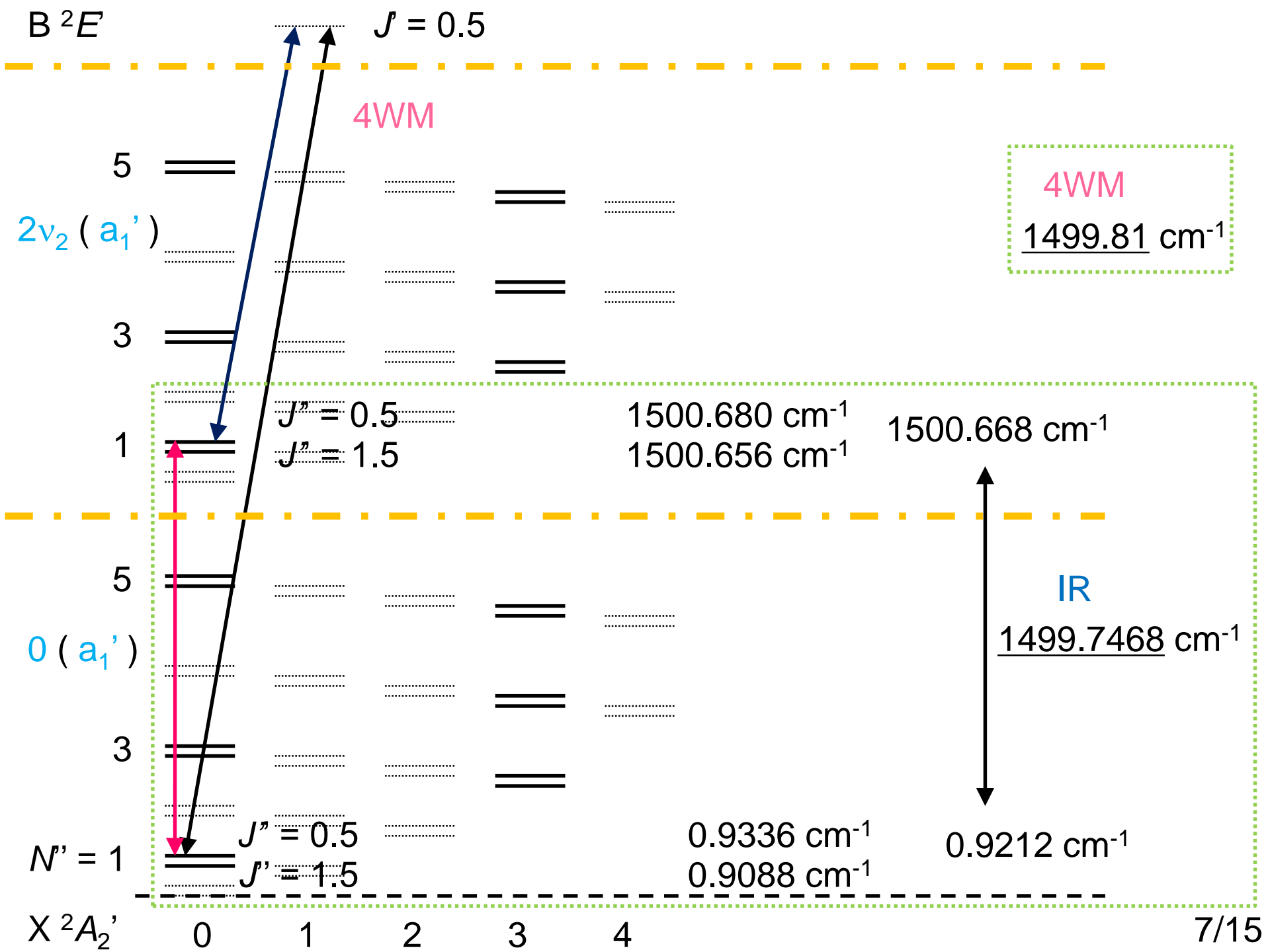


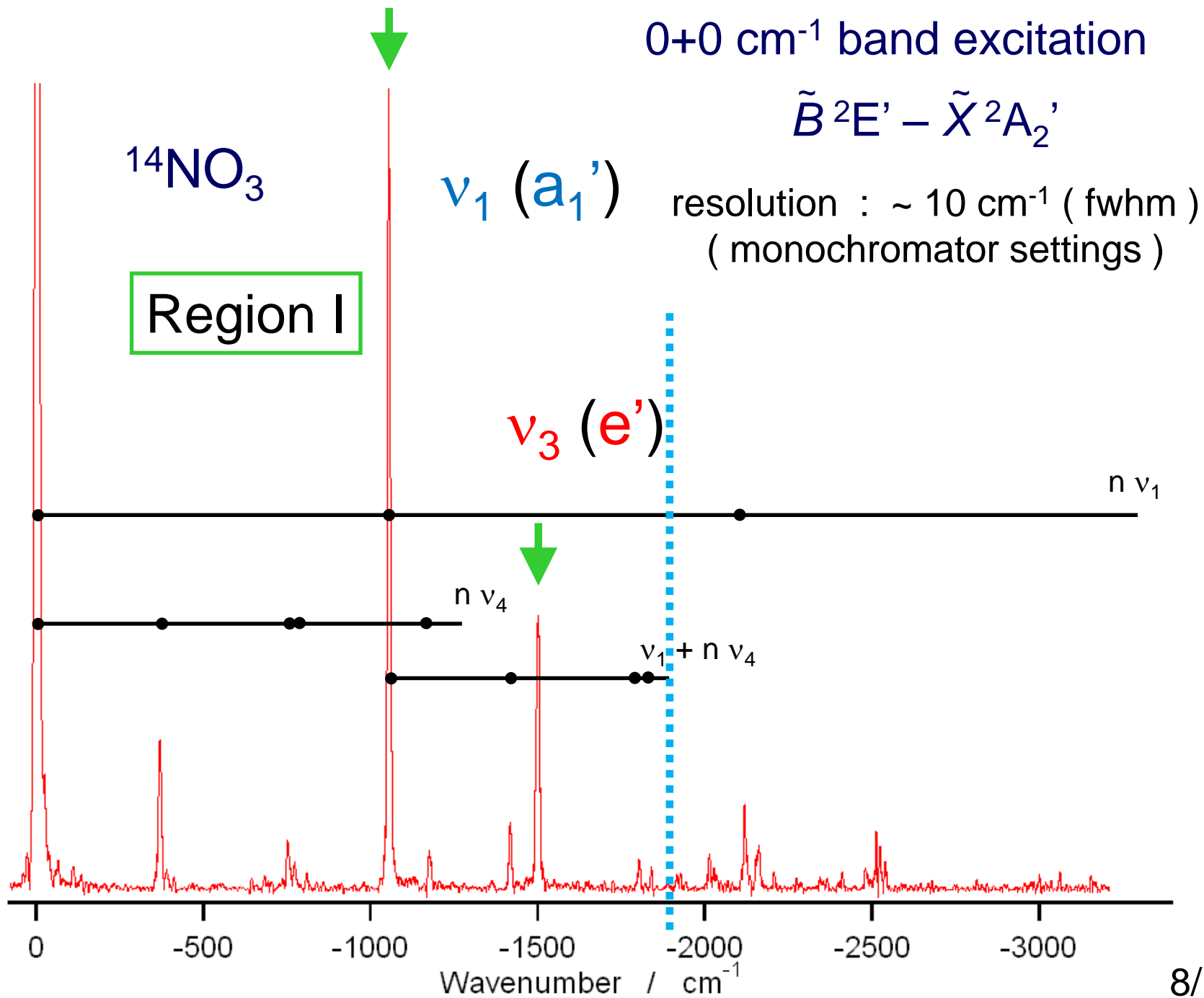
LIF Int. / arb.

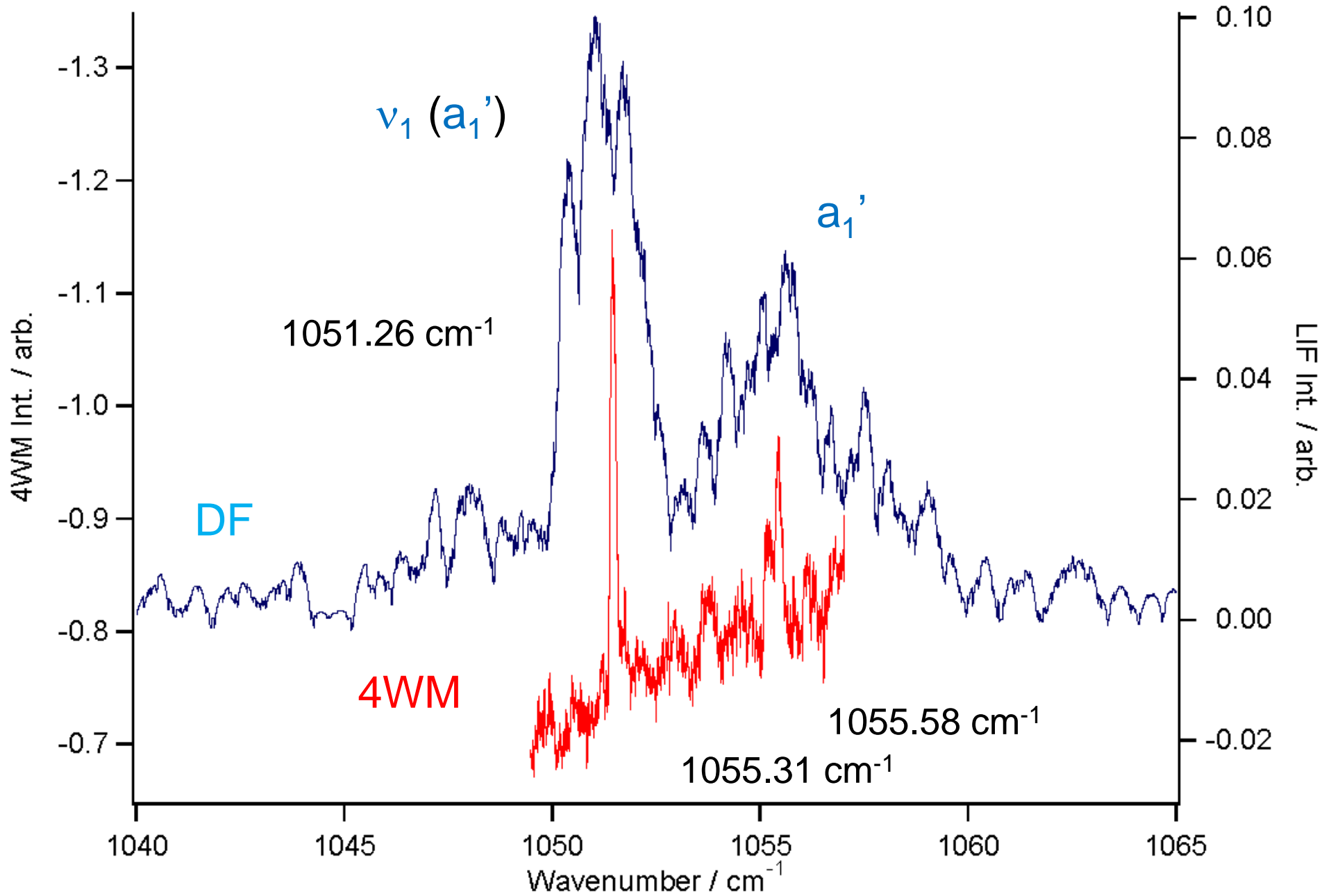


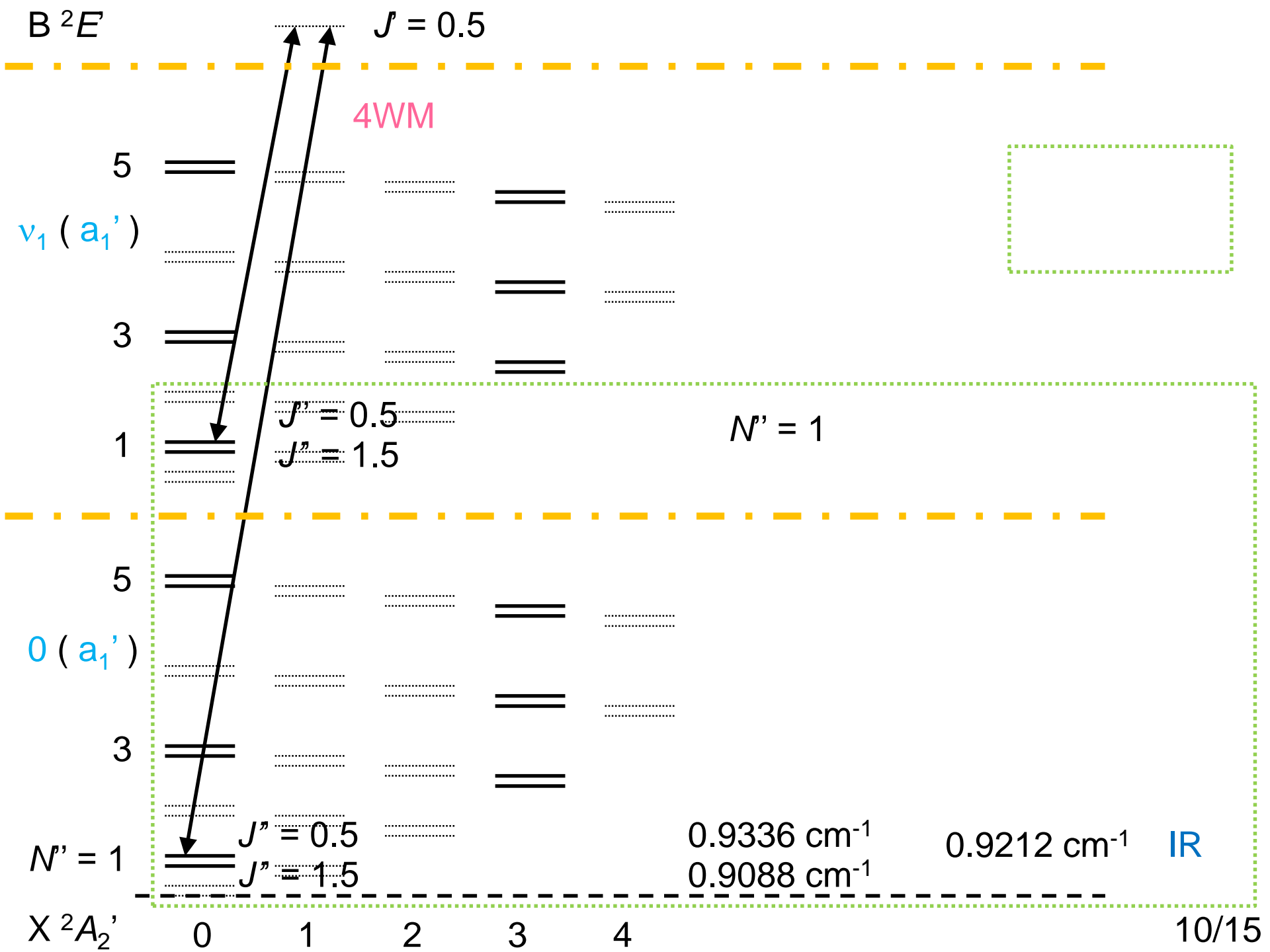


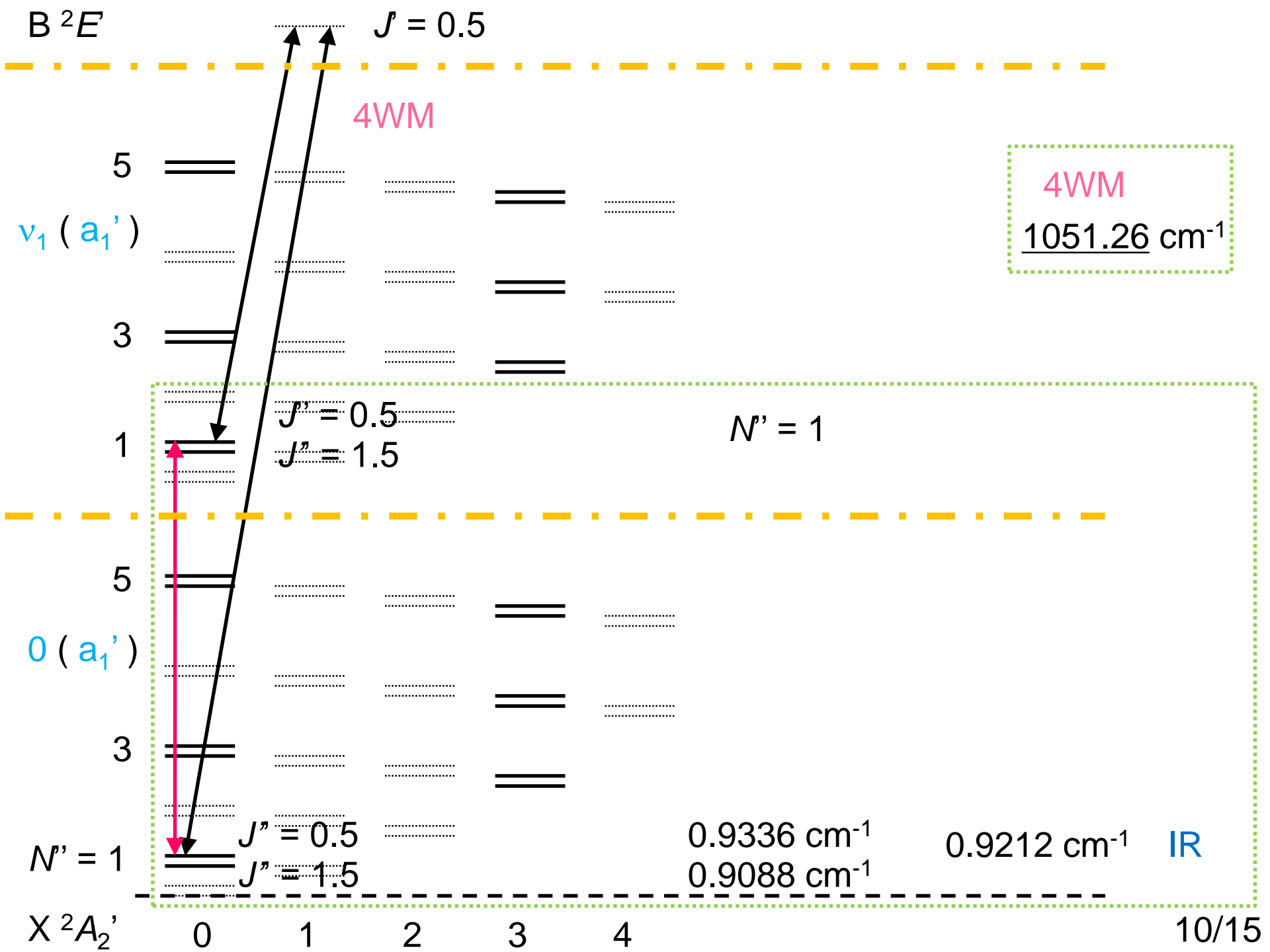


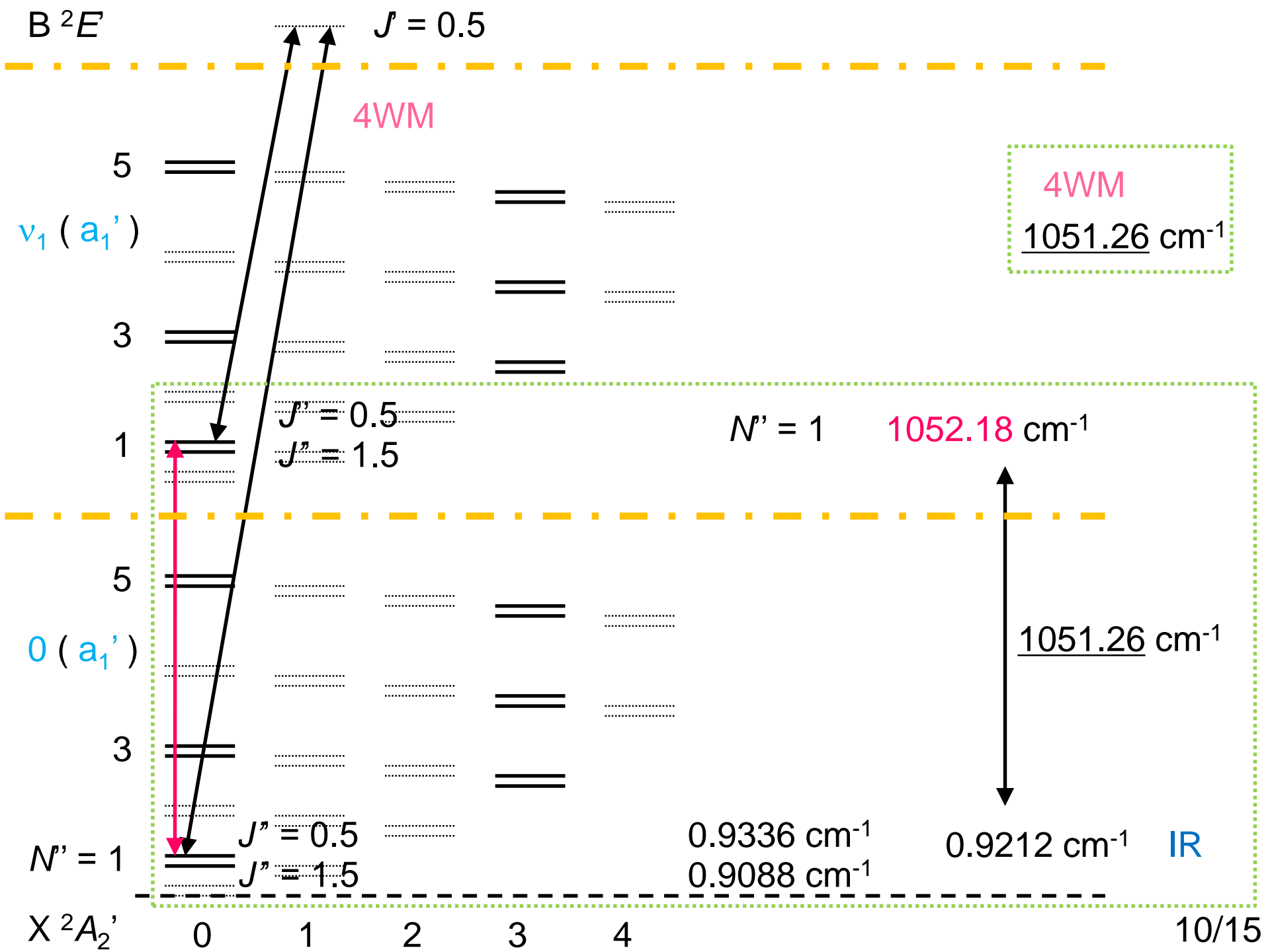


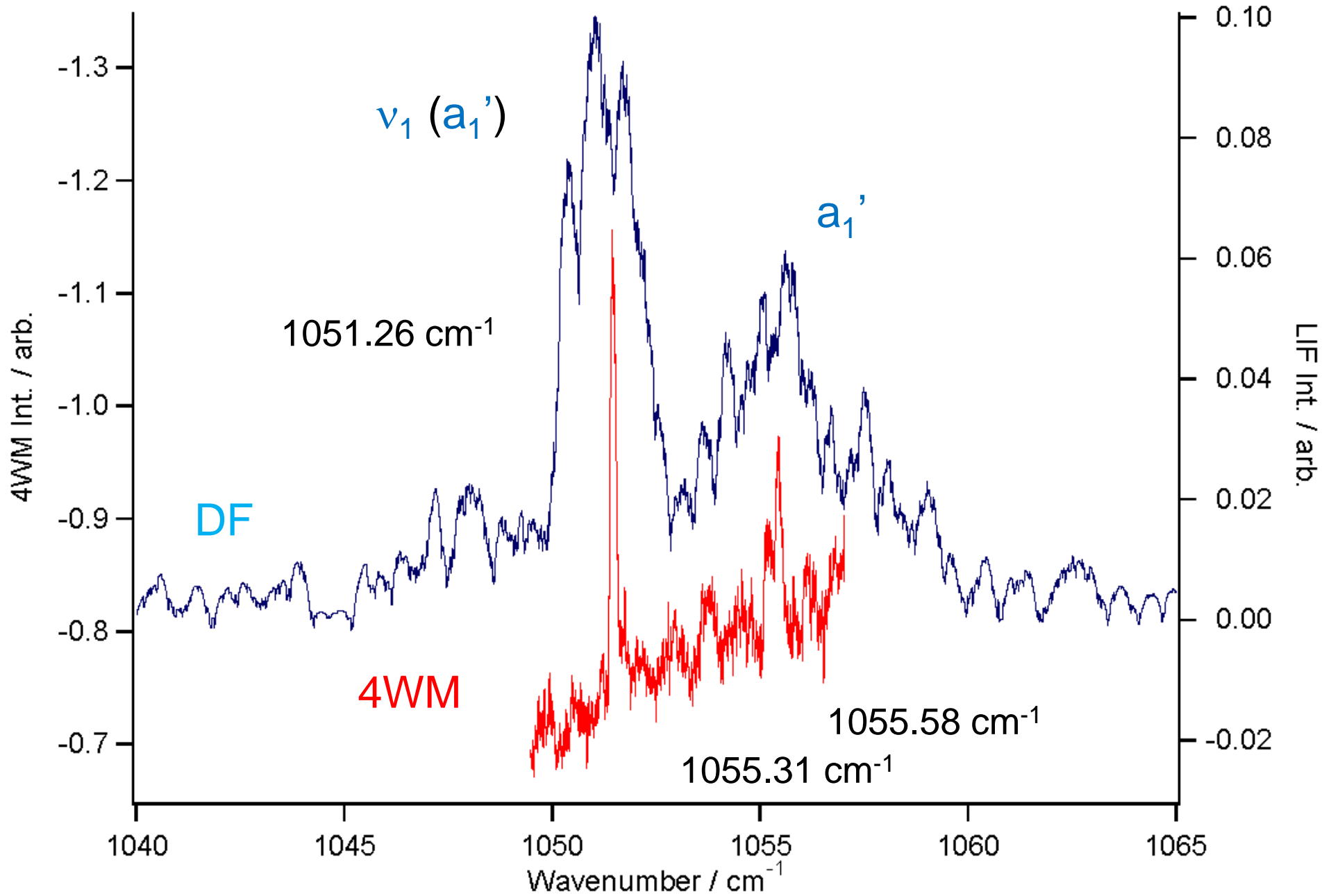




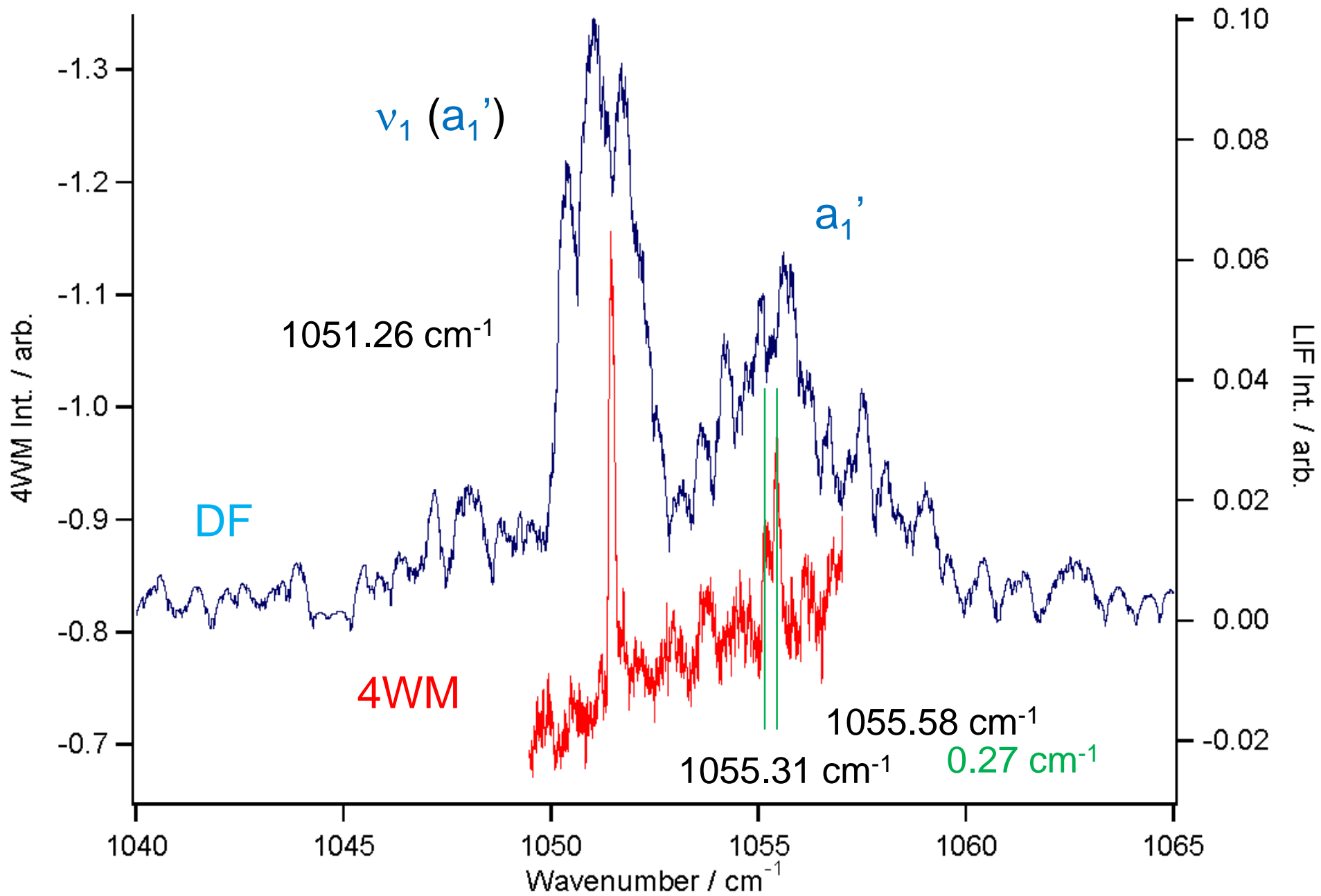


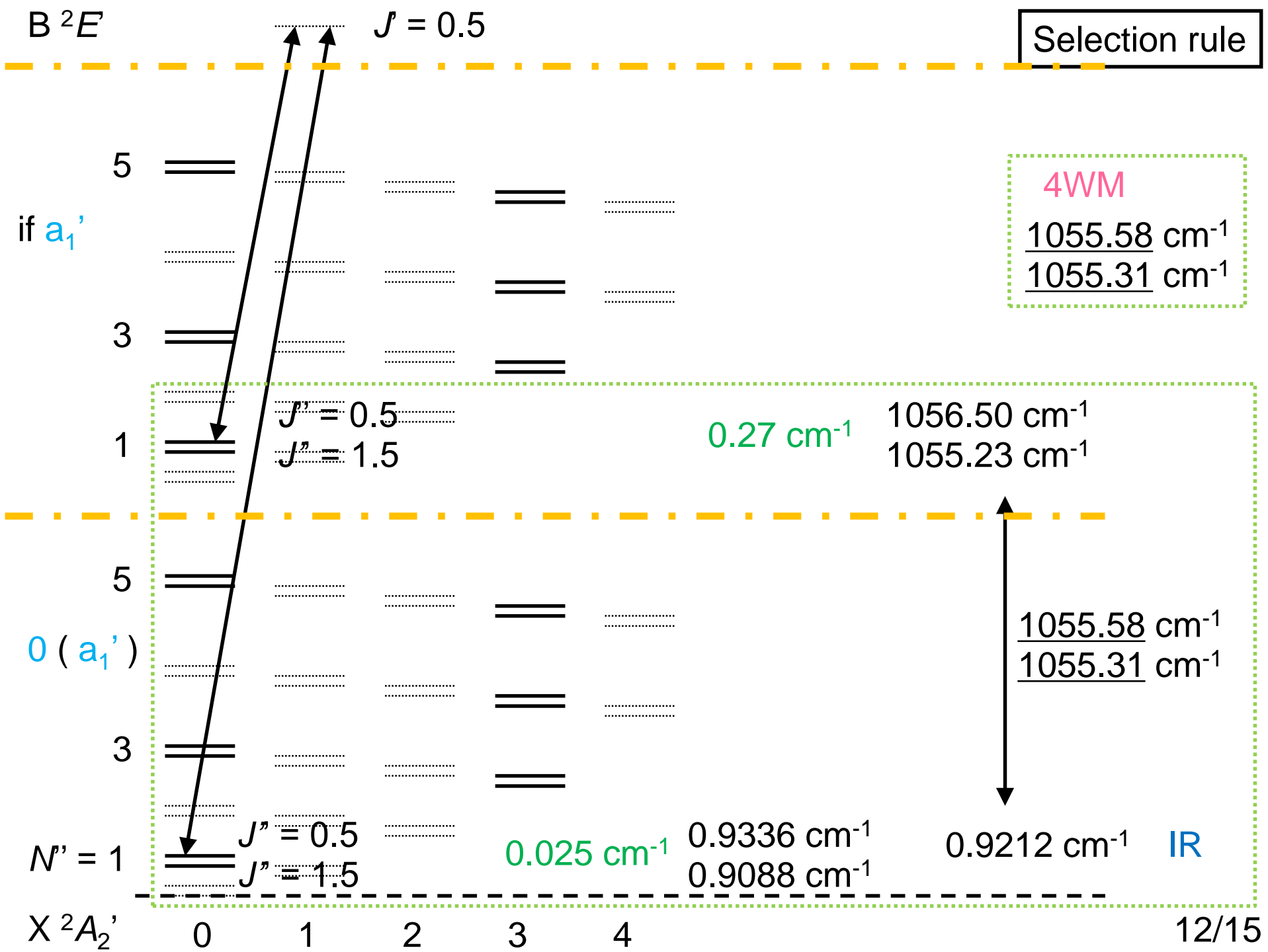


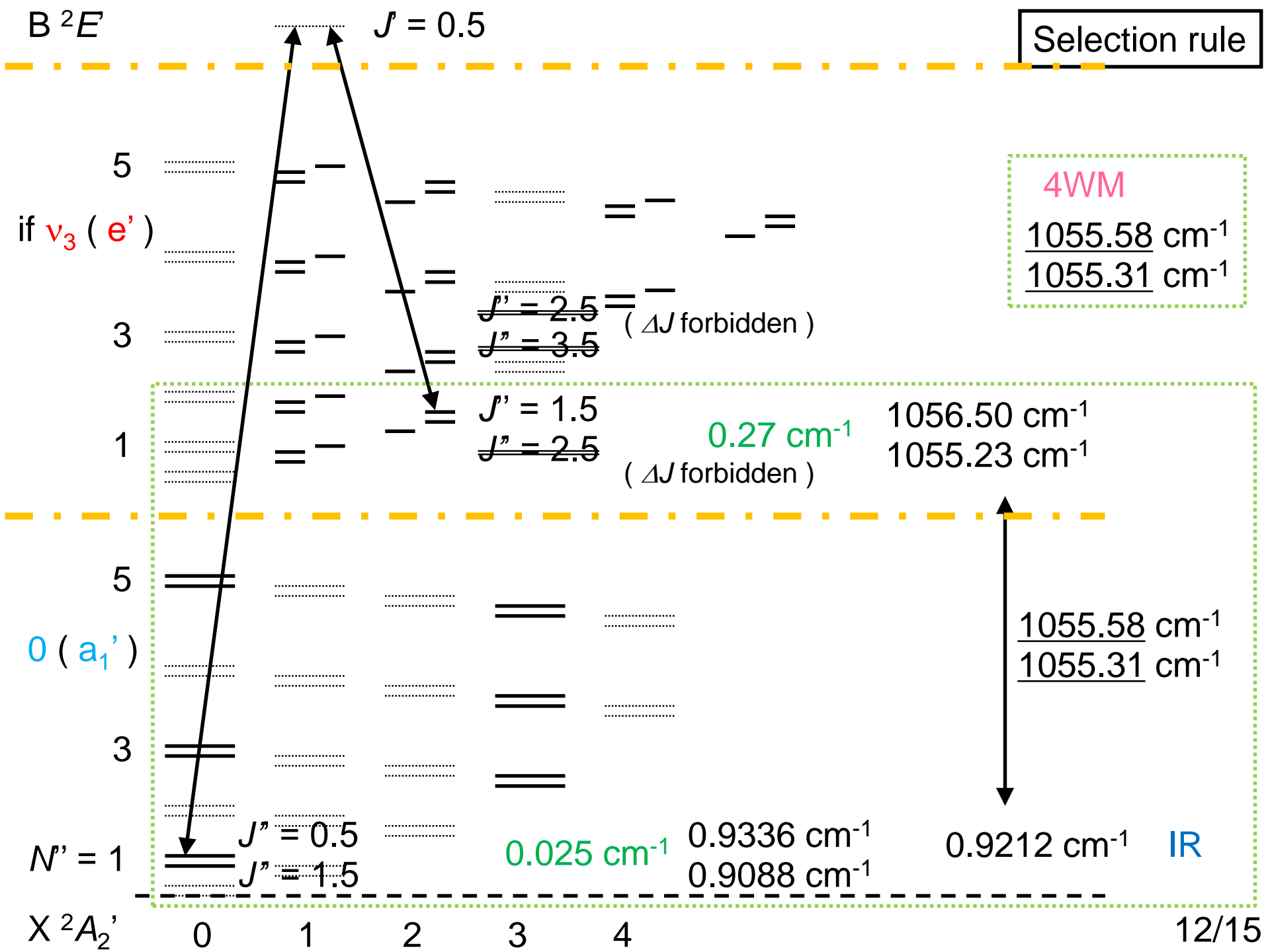


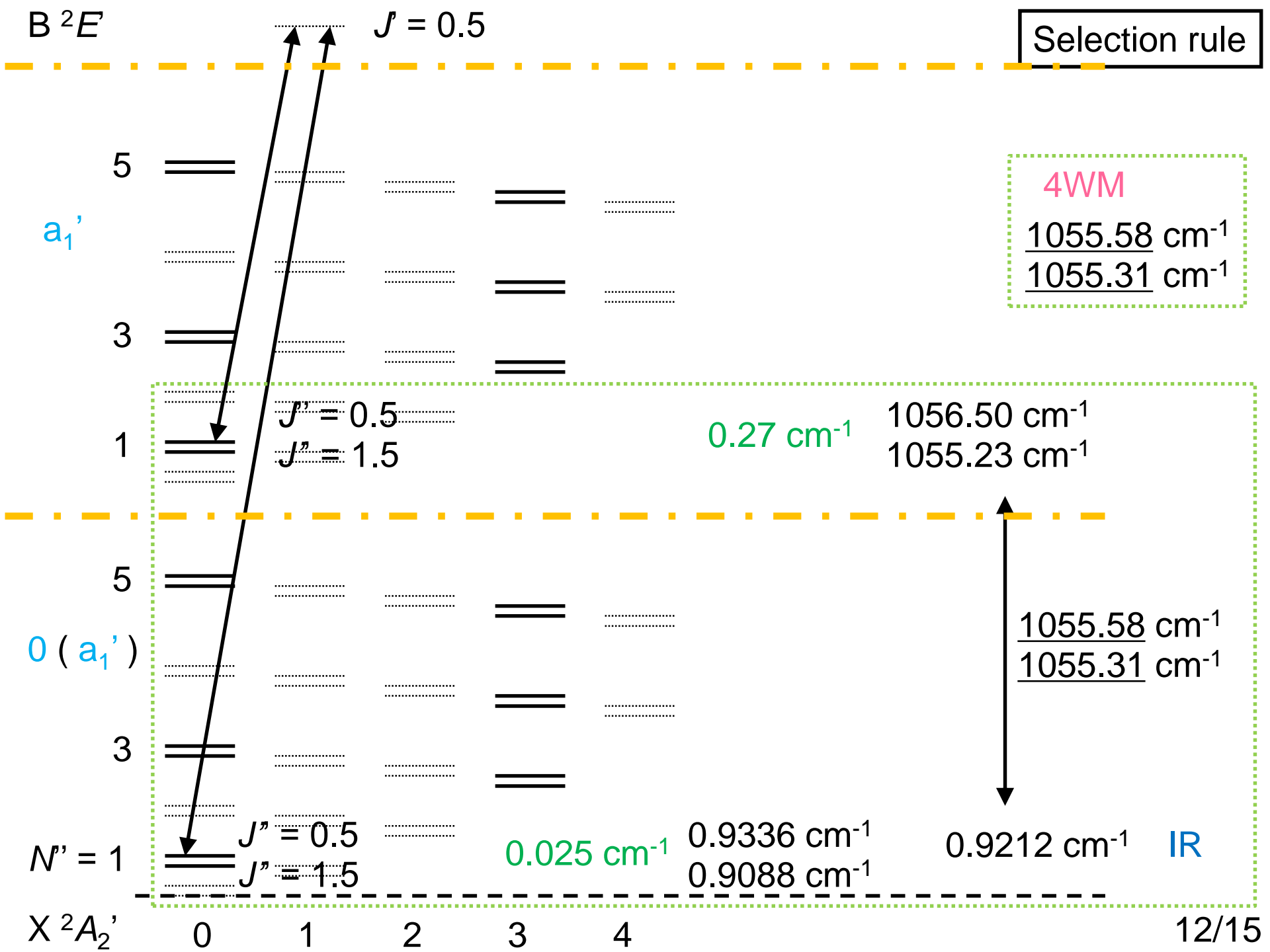












The splitting,  $0.27 \text{ cm}^{-1}$ , at  $N = 1$  of the new  $a_1'$  vibronic level.

We think the new  $a_1'$  level is assigned to  $3 \nu_4 (a_1')$ .

We consider two limiting cases;

(1)  $l$  and  $\Sigma$  are independent;

$$|\nu_4, l, \Sigma\rangle = |\nu_4, l\rangle |\Sigma\rangle$$

(2)  $l$  and  $\Sigma$  are coupled;

$$|\nu_4, l, \Sigma, P\rangle \neq |\nu_4, l\rangle |\Sigma\rangle$$

$$(P = l + \Sigma)$$

The splitting,  $0.27 \text{ cm}^{-1}$ , at  $N = 1$  of the new  $a_1'$  vibronic level.

$a_1'$  ( $l = 0$ ) vibronic levels  
(zero-vibrational,  $v_1$  fundamental)

$$|v_4, l, \Sigma\rangle = |v_4, l\rangle |\Sigma\rangle$$

$$N = 1 \quad \begin{array}{l} \underline{\underline{J = 0.5}} \quad |0, 0\rangle |-\frac{1}{2}\rangle \\ \underline{\underline{J' = 1.5}} \quad |0, 0\rangle |+\frac{1}{2}\rangle \\ K = 1 \end{array}$$

splitting : spin-rotation ( $0.025 \text{ cm}^{-1}$ )

$a_1'$  ( $l = \pm 3$ ) vibronic levels  
 $3 v_4 (a_1')$

$$|v_4, l, \Sigma\rangle = |v_4, l\rangle |\Sigma\rangle$$

$$N = 1 \quad \begin{array}{l} \underline{\underline{J = 0.5}} \quad |3, 0; a_1'\rangle |-\frac{1}{2}\rangle \\ \underline{\underline{J = 1.5}} \quad |3, 0; a_1'\rangle |+\frac{1}{2}\rangle \\ K = 1 \end{array}$$

$$\begin{array}{l} a_1' : |3, +3\rangle + |3, -3\rangle : |3, 0; a_1'\rangle \\ a_2' : |3, +3\rangle - |3, -3\rangle : |3, 0; a_2'\rangle \end{array}$$

The splitting,  $0.27 \text{ cm}^{-1}$ , at  $N = 1$  of the new  $a_1'$  vibronic level.

$a_1'$  ( $l = 0$ ) vibronic levels  
(zero-vibrational,  $v_1$  fundamental)

$$|v_4, l, \Sigma\rangle = |v_4, l\rangle |\Sigma\rangle$$

$$N = 1 \quad \begin{array}{l} J = 0.5 \quad |0, 0\rangle |-\frac{1}{2}\rangle \\ \hline J' = 1.5 \quad |0, 0\rangle |+\frac{1}{2}\rangle \\ K = 1 \end{array}$$

splitting : spin-rotation ( $0.025 \text{ cm}^{-1}$ )

$a_1'$  ( $l = \pm 3$ ) vibronic levels  
 $3 v_4 (a_1')$

$$|v_4, l, \Sigma\rangle = |v_4, l\rangle |\Sigma\rangle$$

$$N = 1 \quad \begin{array}{l} J = 0.5 \quad |3, 0; a_1'\rangle |-\frac{1}{2}\rangle \\ \hline J = 1.5 \quad |3, 0; a_1'\rangle |+\frac{1}{2}\rangle \\ K = 1 \end{array}$$

$$|v_4, l, \Sigma, P\rangle \neq |v_4, l\rangle |\Sigma\rangle \\ (P = l + \Sigma)$$

$$N = 1 \quad \begin{array}{l} J = 0.5 \quad |0, 0, -\frac{1}{2}, -\frac{1}{2}\rangle \\ \hline J = 1.5 \quad |0, 0, +\frac{1}{2}, +\frac{1}{2}\rangle \\ K = 1 \end{array}$$

$$|v_4, l, \Sigma, P\rangle \neq |v_4, l\rangle |\Sigma\rangle \\ (P = l + \Sigma)$$

$$N = 1 \quad \begin{array}{l} J = 0.5 \quad |3, 0, \mp\frac{1}{2}, \mp\frac{1}{2}\rangle \\ \hline J = 1.5 \quad |3, 0, \pm\frac{1}{2}, \pm\frac{1}{2}\rangle \\ K = 1 \end{array}$$

$$a_1' : |3, 0, \pm\frac{1}{2}, \pm\frac{1}{2}\rangle = |3, +3, +\frac{1}{2}, +\frac{7}{2}\rangle + |3, -3, -\frac{1}{2}, -\frac{7}{2}\rangle$$

$$a_1' : |3, 0, \mp\frac{1}{2}, \mp\frac{1}{2}\rangle = |3, +3, -\frac{1}{2}, +\frac{5}{2}\rangle + |3, -3, +\frac{1}{2}, -\frac{5}{2}\rangle$$

The splitting,  $0.27 \text{ cm}^{-1}$ , at  $N = 1$  of the new  $a_1'$  vibronic level.

$a_1'$  ( $l = 0$ ) vibronic levels  
(zero-vibrational,  $v_1$  fundamental)

$$|v_4, l, \Sigma\rangle = |v_4, l\rangle |\Sigma\rangle$$

$$N = 1 \quad \begin{array}{l} \underline{\underline{J = 0.5}} \quad |0, 0\rangle |-\frac{1}{2}\rangle \\ \underline{\underline{J' = 1.5}} \quad |0, 0\rangle |+\frac{1}{2}\rangle \\ K = 1 \end{array}$$

splitting : spin-rotation ( $0.025 \text{ cm}^{-1}$ )

$$|v_4, l, \Sigma, P\rangle \neq |v_4, l\rangle |\Sigma\rangle$$

( $P = l + \Sigma$ )

$$N = 1 \quad \begin{array}{l} \underline{\underline{J = 0.5}} \quad |0, 0, -\frac{1}{2}, -\frac{1}{2}\rangle \\ \underline{\underline{J = 1.5}} \quad |0, 0, +\frac{1}{2}, +\frac{1}{2}\rangle \\ K = 1 \end{array}$$

$a_1'$  ( $l = \pm 3$ ) vibronic levels  
 $3 v_4 (a_1')$

$$|v_4, l, \Sigma\rangle = |v_4, l\rangle |\Sigma\rangle$$

$$N = 1 \quad \begin{array}{l} \underline{\underline{J = 0.5}} \quad |3, 0; a_1'\rangle |-\frac{1}{2}\rangle \\ \underline{\underline{J = 1.5}} \quad |3, 0; a_1'\rangle |+\frac{1}{2}\rangle \\ K = 1 \end{array}$$

Different splitting is expected.

(**S**-uncoupling among the  $P$  bases ?)

$$|v_4, l, \Sigma, P\rangle \neq |v_4, l\rangle |\Sigma\rangle$$

( $P = l + \Sigma$ )

$$N = 1 \quad \begin{array}{l} \underline{\underline{J = 0.5}} \quad |3, 0, \mp\frac{1}{2}, \mp\frac{1}{2}\rangle \\ \underline{\underline{J = 1.5}} \quad |3, 0, \pm\frac{1}{2}, \pm\frac{1}{2}\rangle \\ K = 1 \end{array}$$

$$a_1' : |3, 0, \pm\frac{1}{2}, \pm\frac{1}{2}\rangle = |3, +3, +\frac{1}{2}, +\frac{7}{2}\rangle + |3, -3, -\frac{1}{2}, -\frac{7}{2}\rangle$$

$$a_1' : |3, 0, \mp\frac{1}{2}, \mp\frac{1}{2}\rangle = |3, +3, -\frac{1}{2}, +\frac{5}{2}\rangle + |3, -3, +\frac{1}{2}, -\frac{5}{2}\rangle$$



## Conclusions

LIF DF (higher resolution) and 2C-R4WM spectra have been measured for vibronic bands at the 1050 and 1500  $\text{cm}^{-1}$  regions, and three  $a_1'$  bands,  $\nu_1$ ,  $2\nu_2$ , and  $3\nu_4$ , give the 2C-R4WM signals.

- The  $2\nu_2$  ( $a_1'$ ) spectrum confirms that 2C-R4WM can be applied to characterize vibrational levels.
- The term value of  $\nu_1$  has been determined, though the spin splitting is not resolved.
- 2C-R4WM spectrum of the new 1055  $\text{cm}^{-1}$  band shows larger spin splitting at  $K = 1$  and  $N = 1$ , and this indicates this is an  $a_1'$  band, and this agrees with our  $3\nu_4$  ( $a_1'$ ) assignment of this band.

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## Problems of computational results

### 1000 $\text{cm}^{-1}$ ( near $\nu_1$ ) region

- **The  $\nu_3$  fundamental**, which predicted to be at  $\sim 1000 \text{ cm}^{-1}$  ( near  $\nu_1$  ), **has not been observed** yet by IR spectroscopy.

Stanton (2007) :            The  $\nu_3$  fundamental is too weak.

Viel & Einfeld (2018) :    It should be observed.

### 1500 $\text{cm}^{-1}$ ( near “traditional” $\nu_3$ ) region

- **The  $a_2'$  level**, which assigned to be the  $\nu_3+\nu_4$  combination, **has not been observed yet** by IR spectroscopy.

Stanton (2007) :            The  $a_3'$  level should be observed.

Viel & Einfeld (2018) :    It should be observed.

Symmetric molecules, like  $\text{NO}_3$ , have strong  $\nu_3$  fundamental, as “traditional”  $\nu_3$  assignment for  $\text{NO}_3$ .