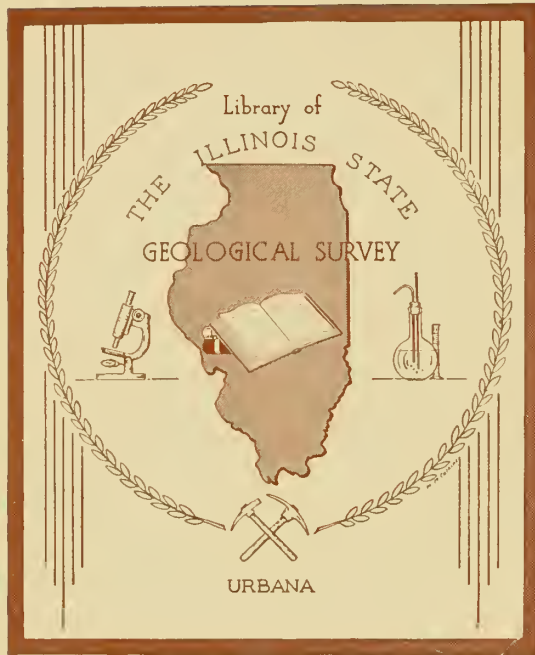





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AROMATIC FLUORINE COMPOUNDS  
I. THE SYNTHESIS OF 2,5- AND 3,5-DIFLUORO-  
BENZOTRIFLUORIDES

By

G. C. FINGER AND F. H. REED

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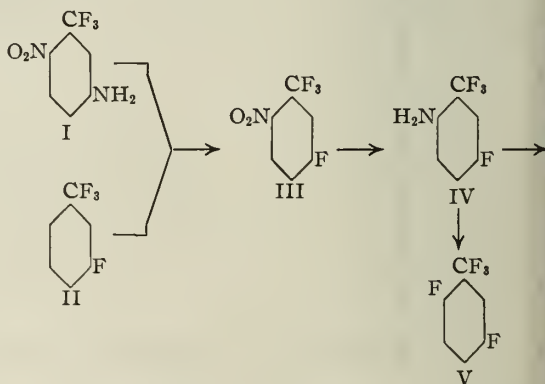
[CONTRIBUTION FROM THE GEOCHEMICAL SECTION OF THE ILLINOIS STATE GEOLOGICAL SURVEY]

## Aromatic Fluorine Compounds. I. The Synthesis of 2,5- and 3,5-Difluorobenzotrifluorides<sup>1</sup>

BY G. C. FINGER<sup>2</sup> AND F. H. REED<sup>3</sup>

As part of a study in this Laboratory, of the synthesis and properties of aromatic fluorine compounds, 2,5- and 3,5-difluorobenzotrifluoride have been synthesized.

A Schiemann reaction<sup>4</sup> on 2-nitro-5-aminobenzotrifluoride (I) gave 2-nitro-5-fluorobenzotrifluoride (III), but this compound was found to be more conveniently prepared by the nitration of 3-fluorobenzotrifluoride (II); the structure of the nitro compound in this synthesis was established by the identity of the acetyl derivatives at



(1) Presented before the Organic Division at the 107th Meeting of the American Chemical Society, Cleveland, Ohio, April 5, 1944. Published with the permission of the Chief of the Illinois State Geological Survey.

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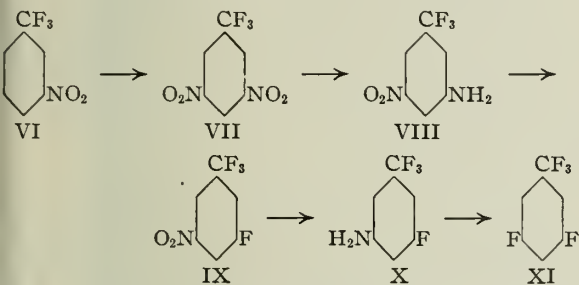
(4) Schiemann and Winkelmüller, *Ber.*, **60**, 1186 (1927).

(IV). An iron reduction<sup>5</sup> of III gave a quantitative yield.

(5) Lukasevich and Vorshilova, *Compt. rend. acad. sci. U. R. S. S.*, **2**, 344 (1935).

yield of 2-amino-5-fluorobenzotrifluoride (IV). This amine was converted to 2,5-difluorobenzotrifluoride (V) by the Schiemann reaction.<sup>6</sup>

A fuming nitric-sulfuric acid mixture with 3-nitrobenzotrifluoride (VI) gave a 50% yield of 3,5-dinitrobenzotrifluoride (VII). By means of the Murray and Waters<sup>7</sup> ammonium sulfide method, the dinitro compound was reduced to 3-amino-5-nitrobenzotrifluoride (VIII). The sub-



sequent compounds were prepared by the same general reactions as applied to I, III, and IV, respectively.

The boiling points of 2,5- and 3,5-difluorobenzotrifluoride, 107–108 and 94–95°, respectively, are not appreciably different from the less fluorinated compounds, toluene, b. p. 110.8; benzotrifluoride, b. p. 102.3; and *m*-fluorobenzotrifluoride, b. p. 100.9, in spite of the great difference in their molecular weights.

### Experimental<sup>8,9</sup>

**2-Nitro-5-fluorobenzotrifluoride (III).**—This compound was prepared by two methods; (1) the Schiemann reaction on 2-nitro-5-aminobenzotrifluoride (I), and (2) the nitration of 3-fluorobenzotrifluoride (II).

Approximately 15 g. of 2-nitro-5-aminobenzotrifluoride<sup>10</sup> in 75 cc. of concd. hydrochloric acid was diazotized with a solution of 6.9 g. of sodium nitrite in 15 cc. of water at 0°, and finally diluted with 25 cc. of water. The diazonium fluoborate was obtained by adding a solution of 40 g. of sodium fluoborate in 75 cc. of water to the filtered diazonium chloride. After recovery by filtration, the insoluble diazonium fluoborate was dried and 21 g. or 94% of the theoretical yield was obtained. After a thermal decomposition of the diazonium fluoborate in the usual manner, the crude product was isolated. A yield of 6 g., 41% based on the diazonium fluoborate, or 39% based on the amine, was obtained. Pure 2-nitro-5-fluorobenzotrifluoride is a heavy almost colorless liquid with a mild nitrobenzene odor; b. p. 198–199°.

*Anal.* Calcd. for  $C_7H_5O_2NF_4$ : C, 40.21; H, 1.45; N, 6.70. Found: C, 40.22; H, 1.49; N, 6.95.

The second method is a "mixed acid" nitration of 3-fluorobenzotrifluoride. To a stirred mixture of 105 g. of concd. nitric acid and 336 g. of concd. sulfuric acid, 164 g. of 3-fluorobenzotrifluoride (II)<sup>11,12</sup> was added at such a rate that the temperature was maintained at 30–35°. At the end, the temperature was raised to 60° for fifteen minutes.

(6) In all cases where the Schiemann reaction was used, a solution of sodium fluoborate was used in place of fluoboric acid.

(7) Murray and Waters, *THIS JOURNAL*, **60**, 2818 (1938).

(8) All melting points are uncorrected.

(9) The authors are indebted to Mr. H. S. Clark, microanalyst, for the analyses given in this investigation.

(10) Rouche, *Bull. classe sci., Acad. roy. Belg.*, 346 (1927).

(11) Booth, Elsey and Burchfield, *THIS JOURNAL*, **57**, 2066 (1935).

(12) Aelony, *ibid.*, **56**, 2063 (1934).

After the removal of the lower "spent acid" layer, the crude product was washed successively with water, 2% sodium carbonate solution, and water, and dried with anhydrous magnesium sulfate. A yield of 188 g., or 90%, of crude 2-nitro-5-fluorobenzotrifluoride was obtained. This material was sufficiently pure for the subsequent reduction.

In order to establish the composition of the crude nitration product, it was reduced with iron to the crude amine, and the resulting amine was acetylated with a mixture of acetic anhydride in acetic acid. Recrystallization of the crude acetyl derivative from carbon tetrachloride, along with the concentration of the mother liquor, gave only one compound, namely, the acetyl derivative of 2-amino-5-fluorobenzotrifluoride. As no other acetyl derivative or readily isolable compound was obtained it was concluded that the nitration of 3-fluorobenzotrifluoride with "mixed acid" gives almost exclusively 2-nitro-5-fluorobenzotrifluoride.

**2-Amino-5-fluorobenzotrifluoride (IV).**—To a stirred, refluxing mixture of 272 g. of iron filings in 375 cc. of ammonium chloride (0.78 *N*), 313.5 g. of 2-nitro-5-fluorobenzotrifluoride was added slowly. The reduction is exothermic and the rate of refluxing is readily controlled by the addition of the nitro compound. Refluxing and stirring were continued for three hours after the nitro compound had been added. Steam distillation removed the amine as a colorless, heavy liquid, giving a yield of 256 g. or 96%. The pure amine has a mild aniline odor, b. p. 70–72° (17.5 mm.) or 170.5°.

The acetyl derivative was prepared in the usual way, and recrystallization from 50% aqueous ethanol gave white needles, m. p. 121–121.5°. This derivative served as a reference compound in proving the structure of the nitro compound derived in the nitration of 3-fluorobenzotrifluoride.

*Anal.* Calcd. for  $C_7H_7ONF_4$ : C, 48.87; H, 3.19; N, 6.33. Found: C, 48.82; H, 3.32; N, 6.45.

**2,5-Difluorobenzotrifluoride (V).**—A mixture of 89.5 g. of 2-amino-5-fluorobenzotrifluoride in 300 cc. of concd. hydrochloric acid was diazotized at 0° with a solution of 37 g. of sodium nitrite dissolved in 50 cc. of water. To the filtered diazonium solution at 0°, 109.8 g. of sodium fluoborate dissolved in 132 cc. of water was added, and the diazonium fluoborate precipitated as a white product. After removal of the diazonium fluoborate by filtration and drying, a yield of 118 g. or 85% was obtained. It was necessary to use an efficient condensing system for the thermal decomposition of the fluoborate due to the high volatility of 2,5-difluorobenzotrifluoride. The condensing system consisted of two vertical Graham condensers (60 cm. long) cooled with brine at 0–10°, and connected so that the gases and vapors from the decomposition flask passed through one condenser and a receiver to the second condenser and receiver. The receivers were cooled also and the major portion of the product collected in the first receiver. The diazonium fluoborate is quite stable and it is necessary to apply heat continually for a progressive decomposition. After the crude product was taken up in ether, the ether extract was made faintly alkaline with dilute sodium hydroxide, and steam distilled. As the difluoro compound is also volatile with steam, it collected as an ether extract in the distillate. After the ether extract was dried, the ether and the difluoro compound were separated by distillation through a 24-inch vacuum-jacketed column. A yield of 52 g. or 52% of 2,5-difluorobenzotrifluoride, b. p. 107–108°, based on the amine or 67% based on the diazonium fluoborate was obtained.

A treatment with cold concd. sulfuric acid improved the odor of the product. It has a very mild and sweet odor as compared to benzotrifluoride.

Additional physical constants are: f. p. *ca.* –14°,  $n_D^{20}$  1.3942; sp. gr. 1.4244 (15°).

*Anal.* Calcd. for  $C_7H_3F_6$ : C, 46.17; H, 1.66. Found: C, 46.20; H, 1.69.

**3,5-Dinitrobenzotrifluoride (VII).**—To a stirred mixture of 450 g. of fuming nitric acid (d. 1.49–1.5) and 900 g. of fuming sulfuric acid (20%  $SO_3$ ) at 95–100°, 57 g. of 3-

nitrobenzotrifluoride<sup>13</sup> was added slowly and the temperature was controlled carefully so as not to exceed 100° for the first ninety minutes. Then the temperature was raised slowly and maintained at 115° for the next four to five hours. The entire reaction mixture was poured over crushed ice with rapid stirring, and the crude product separated as a white solid. It was removed on a large Buchner filter, washed with ice water, and dried. A yield of 43 g. or 52% was obtained. Recrystallization from methanol gave pure 3,5-dinitrobenzotrifluoride as white crystals, m. p. 49–50°.

*Anal.* Calcd. for  $C_7H_3O_2N_2F_3$ : C, 35.6; H, 1.28; N, 11.86. Found: C, 35.61; H, 1.25; N, 11.83.

**3-Amino-5-nitrobenzotrifluoride (VIII).**—To a refluxing solution of 47.2 g. of 3,5-dinitrobenzotrifluoride dissolved in 700 cc. of ethanol, 500 cc. of an alcoholic ammonium sulfide solution was added slowly over a period of thirty minutes, and refluxing was continued for approximately five hours. The ammonium sulfide<sup>7</sup> was prepared by treating 96 g. of crystalline sodium sulfide with 85.6 g. of ammonium chloride in ethanol, and removing the precipitated sodium chloride by filtration. After the reduction was complete, the reaction mixture was concentrated by evaporation to about 200 cc., cooled, filtered, and the precipitate of sulfur and inorganic salts was washed with ethanol to dissolve any residual nitroamine. The filtrate was poured into a large volume of water thus precipitating the nitroamine. After the crude product had been removed by filtration, it was extracted several times with hot, dilute hydrochloric acid. The acid extract was made alkaline and the nitroamine was isolated. A yield of 33 g. or 80% was obtained. Recrystallization from carbon tetrachloride gave pure 3-amino-5-nitrobenzotrifluoride as yellow flakes, m. p. 80–81.5°.

*Anal.* Calcd. for  $C_7H_5O_2N_2F_3$ : C, 40.78; H, 2.44; N, 13.59. Found: C, 40.87; H, 2.46; N, 13.48.

The acetyl derivative was recrystallized from carbon tetrachloride as white needles, m. p. 134–135°.

(13) Finger, Nachtrieb and Reed, *Trans. Illinois State Acad. Sci.*, **31**, 132 (1939).

*Anal.* Calcd. for  $C_9H_7O_2N_2F_3$ : C, 43.53; H, 2.84; N, 11.29. Found: C, 43.69; H, 2.83; N, 11.20.

**3-Fluoro-5-nitrobenzotrifluoride (IX).**—This compound was prepared from 3-amino-5-nitrobenzotrifluoride by the Schiemann reaction. An 88% yield of the diazonium fluoborate was obtained. The thermal decomposition of the diazonium fluoborate gave a 43% yield of 3-fluoro-5-nitrobenzotrifluoride based on the amine. After purification the pure compound boiled at 180.5°.

*Anal.* Calcd. for  $C_7H_3O_2NF_4$ : C, 40.21; H, 1.45; N, 6.70. Found: C, 40.40; H, 1.44; N, 6.53.

**3-Amino-5-fluorobenzotrifluoride (X).**—This amine was prepared from 3-nitro-5-fluorobenzotrifluoride (IX) by the same general synthesis as indicated for IV. A 90% yield of crude 3-amino-5-fluorobenzotrifluoride was obtained. Recrystallization of the acetyl derivative from carbon tetrachloride gave white needles, m. p. 99.5–100.5°.

*Anal.* Calcd. for  $C_9H_7ONF_4$ : C, 48.88; H, 3.19; N, 6.33. Found: C, 49.11; H, 3.10; N, 6.57.

**3,5-Difluorobenzotrifluoride (XI).**—This compound was prepared from 3-amino-5-fluorobenzotrifluoride by the same general procedure as outlined under V. The yield of the diazonium fluoborate was 94%. Thermal decomposition of this compound gave 3,5-difluorobenzotrifluoride. The pure compound gave the following constants: f. p. ca. –40°, b. p. 94–95°,  $n_D^{20}$  1.3873.

*Anal.* Calcd. for  $C_7H_3F_6$ : C, 46.17; H, 1.66. Found: C, 46.10; H, 1.53

## Summary

The preparation of 2,5- and 3,5-difluorobenzotrifluoride and some of their intermediates is described.

3,5-Dinitrobenzotrifluoride was prepared from 3-nitrobenzotrifluoride with a fuming nitric-sulfuric acid mixture.

URBANA, ILLINOIS

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